

Optimization, Structural, Spectroscopic and Magnetic Studies on Stable Akaganeite Nanoparticles via Co-Precipitation Method

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Abstract With a view to achieve stable the ultrafine akaganeite (β -FeO(OH)) particles, they are synthesized by varying the molar concentration of dextrose, $x = 0.2, 0.4, 0.6, 0.8, 1.0M$, as chelating agent in co-precipitation method. The structural properties studied using the powder X-ray diffraction patterns show that an optimum concentration of dextrose for the formation of akaganeite is $0.6M$. The particles possess tetragonal crystal structures and the size of grains is in the order of nanometer (~ 18 to $29nm$). DSC curves explain the chemical reactions and phase transitions undergone by the sample. FT-IR spectra of the nanoparticles of akaganeite show the absorption bands at 491 and $445cm^{-1}$ in all the samples corresponding to the symmetric stretching vibrations of Fe-O bond and its ligands respectively in the octahedral co-ordinations. UV-Visible spectra of the present set of samples clearly exhibit a change in the optical bandgap from 1.97 to $2.14eV$. The magnetic studies could bring out a weak ferrimagnetic behaviour accompanied by superparamagnetism in the akaganeite nanoparticles.

Keywords X-Ray diffraction, FTIR, Co-precipitation, Akaganeite Nanoparticles, Magnetic study, EPR

1. Introduction

The nanoparticles of oxides/oxyhydroxides of iron continue to attract the attention of the Scientists as they possess a wide range of applications[1-3] such as catalytic materials, sorbents, pigments, flocculants, coatings, gas sensors and ion exchangers. They serve as precursors to phases for information storage media. They are used as model systems for the study of fundamental colloid and surface properties of metal oxides.

As on date there are 16 iron oxides/oxyhydroxides[3] found in nature and they are famous magnetic materials. The material under interest, akaganeite, is one of the polymorphs of iron oxyhydroxides[4,5]. The mineral akaganeite occurring naturally acquires a form β -FeOOH. It was collected from Akagane mine, Iwate Prefecture, by Dr. M. Nambu[6] of Sendai University, Japan, who named it first as "Akaganeite". Akaganeite has a tetragonal structure consisting of double chains of edge-shared octahedra that share corners with adjacent chains to form channels running parallel to the c-axis, which can be occupied by chloride ions and water molecules (refer Figure.1)[7]. Thus formed tunnel

structure makes β -FeOOH an especially interesting material in the areas of catalysis, electrode materials, ion exchange and so on[8-10]. Therefore the synthesis of akaganeite has been extensively investigated owing to its wide application. Earlier Schwetmann and Cornell[11] have prepared akaganeite using the hydrolysis of $FeCl_3$ aqueous solutions at moderate temperature with a preferred acidic pH ($pH < 5$) which avoided the formation of more stable neighborly phases such as hematite and goethite. Later Xiao-Min Lin et.al.[12] used lithographic techniques and obtained particles of size smaller than top-down approach. Meng Meng Lin et.al.[13] have successfully used the co-precipitation of inorganic salts with a basic pH ($pH = 9$) to synthesize some magnetic nanoparticles. Sophie Laurent et.al.[14] have studied the co-precipitation method for the synthesis of nanoparticles of some iron oxides and listed out many advantages of using that method. They claimed that it is one of the simplest and most efficient chemical pathways to prepare the iron oxide nanoparticles due to its versatility, relatively low budget, feasibility to scale up, and the hydrophilic surface character of the intermediary products. Despite a broad distribution of particle sizes, this method supports atomic level mixing of individual components thereby reducing the diffusion distances, which results in the nanocrystalline materials at lower temperatures compared to other methods.

The broad size distribution has been overcome by the use of the different organic additives as stabilization and / or

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reducing agents such as Poly Vinyl Alcohol[15,16], dextran[17,18], oleic acid[19-21], Ethylene Diamine Tetra Acetic acid[22] and so on. These organic compounds play a key role in inducing the crystallization processes in the solution; firstly in the form of high-molecular-weight compounds (aggregates) and subsequently their associates of the amphiphilic organic molecules of low molecular weight [23].

The role of the organic ions on the metal oxides / oxyhydroxides to form their ultrafine particles can be rationalized by two competing mechanisms[24] namely, (i) chelation of the metal ions prevents monotonic nucleation and hence limits the size of nuclei which thereby supports the faster formation of particles and (ii) the absorption of additives by the nuclei and growing crystals may inhibit the growth of the particles, which favours the formation of small units.

Recently Mallikarjuna et.al.[25] have done a pilot study of the synthesis of nanoparticles of akaganeite using dextrose in co-precipitation. We have undertaken a systematic investigation to prepare highly stable akaganeite nanoparticles by optimizing the quantity of dextrose during co-precipitation, avoid the formation of other coexisting compounds and characterize it for the structural, spectroscopic and magnetic properties.

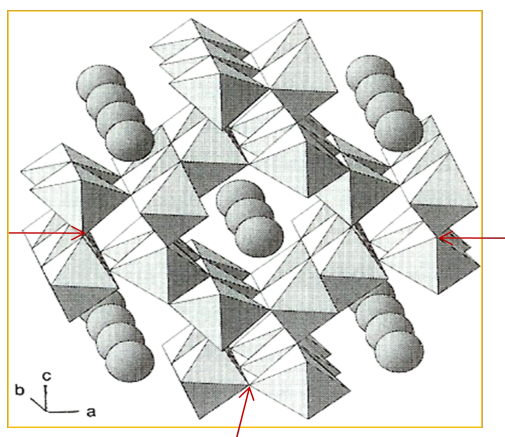


Figure 1. Structure of Akaganeite - Arrangement of octahedral double chains in tunnels with chloride ions in the centre of the tunnels. [Courtesy Ref.7].

2. Experimental Details

2.1. Starting Materials

In this work, the analytical chemical reagents used are ferric chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$) (HIMEDIA RM6353, 99% purity), anhydrous dextrose ($\text{C}_6\text{H}_{12}\text{O}_6$) (HIMEDIA RM077, 99.5% purity) and ammonia solution (NH_3) (sd fine-CHEM Ltd 37140, 25% extra pure). Their respective aqueous solutions are prepared with doubly deionized water.

2.2. Synthesis

Nanoparticles of akaganeite are prepared using the co-precipitation method by mixing the aqueous solutions of iron (III) chloride and dextrose in different molar ratios of

1:x (where, $x = 0.2, 0.4, 0.6, 0.8, 1.0$). The 10M ammonia solution is used as precipitating agent. 1M of ferric chloride hexahydrate and xM of dextrose solutions are mixed in glass beaker and stirred in a magnetic stirrer for 30 minutes to achieve good mixing of the reactants. The solution becomes brownish green in colour with a pH value of 2. The resultant solution in the beaker is then heated to 90°C on a ceramic hot plate, followed by drop wise addition of 10M ammonia by means of a 50mL burette until a pH of 9 was obtained. The addition of ammonia solution and the heating encourage the condensation reaction between adjacent molecules along with the elimination of H_2O and ammonium chloride, organic compounds so as to yield akaganeite. To begin with we observe brownish-yellow colour precipitate in the solution and at the end of the reaction the whole solution becomes orange in colour. The pH value of the solution is found to be ≈ 4 at this stage. When ammonia solution is poured further, a shiny paste is formed and later it settles at the bottom of the beaker as brownish green colour paste. Then the continuation of ammonia addition at $\text{pH} = 9$ and stirring the precipitate for half an hour to prevent agglomeration, the reaction attains a conducive completion. Thus obtained final solution is air cooled to room temperature. The resulting precipitate is filtered and washed with distilled water several times to remove chloride ions. The final colloid or paste like product is poured in a clean silica crucible and heated at 90°C for two hours inside a furnace in order to remove the residual inorganic and organic compounds present in the water soluble colloid. The material taken from the furnace appears like dried clay. It is ground using the agate mortar and pestle to form yellowish-brown colour powder of particles of akaganeite.

2.3. Characterization

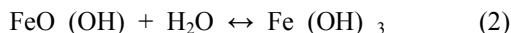
The structure, phase and size of particles of the samples are determined using powder X-Ray Diffractometer, (XRD), PANalytical (Model: X'Pert PRO) equipped with $\text{Cu-K}\alpha$ radiation of wavelength 1.5406 \AA , energy 1000W and Graphite monochromator. Differential Scanning Calorimetric (DSC) study on akaganeite samples is done using TA INSTRUMENTS Q20 SERIES. The Fourier Transform Infra-Red (FTIR) spectra are taken using Shimadzu (Model: 8700) spectrometer. The sample is studied in KBr matrix in the range from 400 to 4000cm^{-1} . The UV-Visible spectra of the samples are obtained from Ocean optics (Model: DH-2000) instrument. The samples are dissolved in methanol and taken in a quartz cuvette for the UV-Vis study. The magnetization as a function of external magnetic field is studied using Vibrating Sample Magnetometer (VSM) (EG & G Model: 155). Electron Paramagnetic Resonance (EPR) of the powder samples is performed using (Model: JEOL JES-TE 100) at X-band frequency 9.38653GHz with microwave power of 1.00mW in the field range 0 – 500mT and 100kHz field modulation to obtain the first derivative EPR spectrum.

2.4. The Proposed Reaction Mechanism of the Formation of Akaganeite

The hydrolysis of Fe^{3+} from ferric chloride solution proceeds as follows:



Akaganeite (β - FeOOH) precipitates normally under low pH. At high pH of the solution, the reaction (1) will be further transformed as



$\text{Fe}(\text{OH})_3$ is amorphous and may slowly transform in to α - FeOOH and Haematite (α - Fe_2O_3)[26].

In the present work, we have tried to prepare akaganeite at high pH. The hydrolysis reaction of $\text{Fe}(\text{III})$ chloride in the presence of dextrose has acidic nature and we can expect that there may be some formation of coordinate compound as reported by Mamata Mohapatra *et.al.*, [22] using $\text{EDTA} + \text{FeCl}_3 \cdot 6\text{H}_2\text{O}$. But due to the presence of carboxyl groups in dextrose, a change of hydrated $[\text{Fe}(\text{H}_2\text{O})_6]$ to β - FeOOH takes place as follows[27],



3. Results and Discussions

3.1. Structural Analysis (Powder-XRD)

The XRD patterns of samples synthesized with different molar ratios of dextrose in the ferric chloride solution are shown in Figure. 2. The XRD patterns show broadened peaks which are typical of nanoparticles as they have a high degree of internal stress and strain[28]. It is noted that for lower quantity of dextrose viz., $x = 0.2$ and 0.4M a secondary phase fitting to Maghemite ($\gamma\text{-Fe}_2\text{O}_3$) for a $2\theta = 30^\circ$ (JCPDS file No: 89-3850) appears in their respective XRD patterns[29]. The rest of the samples viz., $x = 0.6, 0.8$ and 1M possess well resolved diffraction peaks that match well with the standard akaganeite pattern (JCPDS file No: 34-1266) for *tetragonal* body-centered crystal structure with the space group $I4/m$ [22]. It is interesting to note that the concentration of dextrose has played an important role in acquiring the single phase nanoparticles of akaganeite.

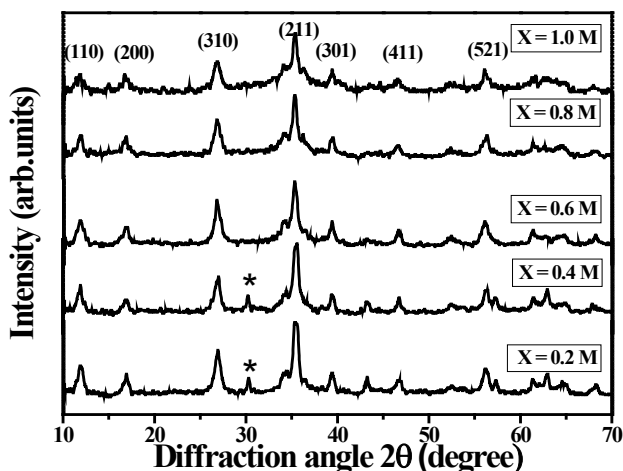


Figure 2. XRD patterns of akaganeite synthesized with different molar ratios of dextrose.

The observed values of lattice parameter and the grain

size of the akaganeite nanoparticles prepared with different molar ratio of dextrose are given in Table 1.

Table 1. Structural parameters of akaganeite obtained from XRD analysis

Molar ratio of dextrose (M)	Lattice parameters (Å)		Grain size (nm)
	a	c	
0.2	10.1490	3.0009	26
0.4	10.1233	2.9960	20
0.6	10.2923	3.0280	29
0.8	10.2203	3.0181	19
1.0	10.1184	3.0009	18

It is seen from Table 1 that the values of lattice parameters change randomly. But it is interesting to note that the grain size of the samples vary in inversely with the lattice parameters for the last three molar ratios. It is seen that 1.0M of dextrose concentration yields particles as small as 18 nm which is well within the range meant for superparamagnetic domain formation for iron oxides[30].

Gash *et.al*[9] have reported that nucleation and growth of akaganeite particles occur at pH values as low as 2. Akaganeite formed for $\text{pH} \geq 3$ is either accompanied with secondary phases like maghemite or transformed into haematite. They also reported that while maintaining low pH, hydrolysis of FeCl_3 solution slows down and unavoidably supports other phase formation. In this work we have prepared akaganeite at comparatively high pH value equal to 9 using dextrose as chelating agent. The role of dextrose is primarily to form a co-ordination complex with Fe^{3+} ion which inhibits the speed of the hydrolysis reaction. Therefore it can be explained that at low concentrations of dextrose viz., 0.2 and 0.4M the complex formation is not complete due to inadequate amount of dextrose compared to Fe^{3+} ions. Thus, appreciable quantity of Fe^{3+} ions forms complex with dextrose and results in akaganeite. Some shallow part of molecules of FeCl_3 which is unable to make complex may undergo hydrolysis rapidly and form maghemite phase. It is explained that for lower concentrations of dextrose akaganeite forms as a major phase along with maghemite as a secondary phase.

The XRD patterns of $x = 0.6, 0.8$ and 1.0M show very distinctly the formation of single phase akaganeite particles. It is understood that at the higher concentrations the dextrose may sufficiently compensate the Fe^{3+} ions and therefore the formation of complex is complete. Since the formation of secondary phases is restricted the pure single phase akaganeite is yielded. It is certain that dextrose indeed plays a vital role in the formation of stable and single phase of akaganeite at higher pH values. Further it is interesting to note that higher molar percentage of dextrose in the solution during the preparation of akaganeite slices the size of the particles down.

3.2. Thermal Analysis (DSC)

The DSC plots of akaganeite particles prepared using different concentrations of dextrose are shown in Figure. 3. There are three major thermal events which can be observed

in the DSC plots of akaganeite particles. A broad endothermic peak in the temperature range from 80 to 120°C is attributed to the evaporation of water molecules present near the surface of the precipitate[31].

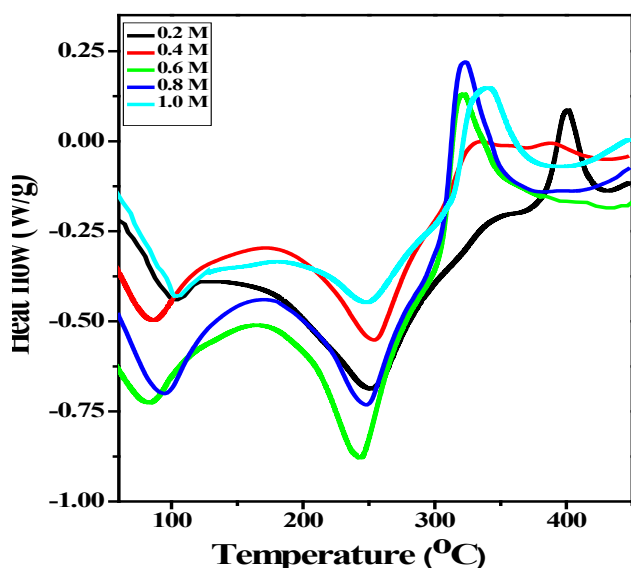


Figure 3. DSC curves of akaganeite samples.

A second endothermic peak with a broad shoulder observed in the temperature range from 240 to 260°C is due to the evaporation of crystalline water molecules. When one moves to the higher temperature a major exothermic peak ranging from 300 to 400°C is indicative of the formation of haematite[32,33]. This process is termed as dehydroxylation i.e., at higher temperatures Cl⁻ ions are expelled from the sample. It corresponds to unlocking of Cl or OH ions from the tunnels of akaganeite structure and onset of haematite structure. It is interesting to note that there is a marked variation in the phase transition temperature between the low and high concentrations of the samples. It is seen that the formation of haematite takes place at temperatures 400 and 370°C for 0.2M and 0.4M samples where as, for 0.6 M, 0.8 M and 1.0 M it happens at 310, 320 and 340°C respectively. Akaganeite, Maghemite and Haematite normally crystallize in tetragonal, cubic with tetragonal super cell, and hexagonal structures respectively. The values of the structural transition temperature and the heat flow associated with the transition are very less for akaganeite when compared to its other polymorphs[34]. Therefore the decrease in the value of transition temperature and heat flow can be solely thought to be dependent on the structure. In the present work the XRD study on akaganeite particles has indicated that they possess tetragonal crystal structure. On that basis, it is concluded that a less amount of energy is required for attaining the tetragonal akaganeite than the hexagonal haematite structures. The possible reason for higher value of transition temperature observed for 0.2 and 0.4M dextrose usage is the presence of haematite or maghemite phase. XRD has also reported a reduction in the size of particles of akaganeite when molar ratio of dextrose is enhanced. It is very much clear that the addition of dextrose speeds up the reaction which supports a faster nucleation and growth of

akaganeite with Cl or OH ions in its tunnels sufficiently. Thus akaganeite prepared is found to be stable and pure as it is formed at low temperature. The excess thermal energy of heat of reaction is absorbed by the surrounding water bed due to higher pH value.

3.3. FT-IR Study

The FT-IR spectra of akaganeite nanoparticles taken in the frequency range of 400- 4000cm⁻¹ are shown in Figure. 4. The IR spectra of the samples show a prominent absorption in the range from 3430 to 3458cm⁻¹ characterizing the O-H stretching vibrational band of water molecule present in akaganeite nanoparticles[35]. This region is very broad and obscured by a large amount of adsorbed water that remains in its structure[36]. Another band located between 1630 and 1648cm⁻¹ is assigned to solitary O-H bending vibration[37]. All the spectra show two very weak absorption peaks near 2925 and 2850cm⁻¹ corresponding to the stretching vibrations of CH₂ and CH₃ respectively. The absorption bands observed at 1350 and 1160cm⁻¹ are due to the vibrations of C=O and C-C bonds respectively. Deliyani et.al[38], have reported that the bands present between 820 and 630 cm⁻¹ are due to the liberation modes of O-H...Cl hydrogen bonds present, which are characteristic of chloride containing akaganeite.

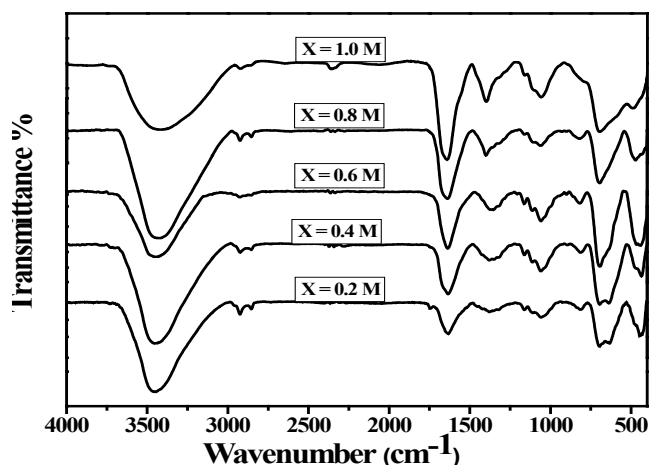


Figure 4. FT-IR spectra of akaganeite samples.

FTIR spectra of all the samples show two persistent absorption bands in the range 430-700cm⁻¹. Kenneth Kauffman et al[36] and Waldron[39] have preferably assigned the bands in the range of frequency from 300 to 700cm⁻¹ to the fundamental vibrations of inorganic ions[40] in the crystal lattice. The absorption band observed from 439 to 491cm⁻¹ in all the samples is assigned to the symmetrical stretching vibrations of Fe-O bond in the octahedral site. Asymmetric stretching vibration of Fe-O-Fe is responsible for the absorption band at 700cm⁻¹[41]. The vibrational modes and force constants of the akaganeite samples with different concentration of dextrose are given in Table 2. It is seen from Table 2 that there is an increase in vibrational frequency and force constants of the octahedral site with the increase in dextrose concentration.

It has been reported[42] that the vibrational frequency

and force constant are inversely proportional to the bond length. Accordingly a decrease in bond length of Fe-O in the octahedral site is well evident. The decrease in lattice constants with the increase in dextrose molar ratio observed in XRD study is supported. Thus we can conclude that the dextrose in the solution plays a role to make Fe^{3+} and O^{2-} ions to go closer. Further the reduction of particle size and enhancement of coupling forces ensure a higher stability to akaganeite nanoparticles prepared in present work.

Table 2. Force constants for corresponding vibrational frequencies of akaganeite.

Concentration of Dextrose (M)	Vibrational frequency of Octahedral site (cm^{-1})	Force constant ($\times 10^3$) N/m
0.2	445	9.7128
0.4	439	9.4527
0.6	451	9.9765
0.8	474	10.8809
1.0	491	11.8246

UV- Visible spectral study

The UV- Visible spectra of akaganeite nanoparticles prepared in the present work show two weak absorption bands in the range 250 -500nm (refer Figure.5.). First one is a broad band in the range from 250 to 380nm corresponding to ligand – metal ion charge transfer transition and second one is weak but broad band found around 460-560nm for

electronic transition between $d \rightarrow d$ bands of iron ions[43].

UV-Visible absorption spectra of akaganeite nanoparticles, can be used to estimate the optical bandgap by applying Tauc model[44,45] in the high frequency absorption region. The plots of $(\alpha h\nu)^2$ with respect to $h\nu$ are shown in Figure. 6 for the data obtained from the respective spectrum. The values of bandgap energy (E_g) are estimated from the graph by extrapolating a tangent line for the linear portion to cut the photon energy axis.

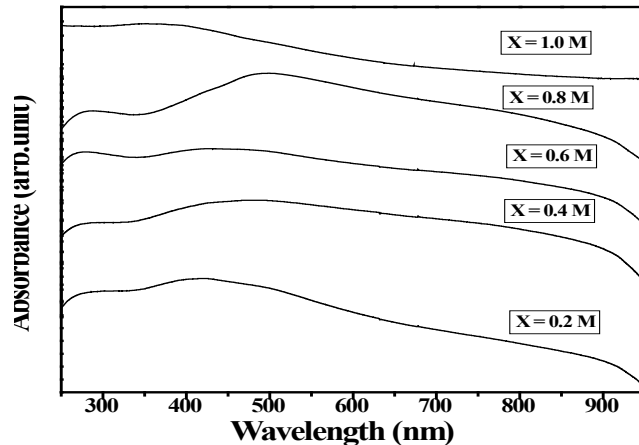


Figure 5. UV- Visible spectra of akaganeite samples.

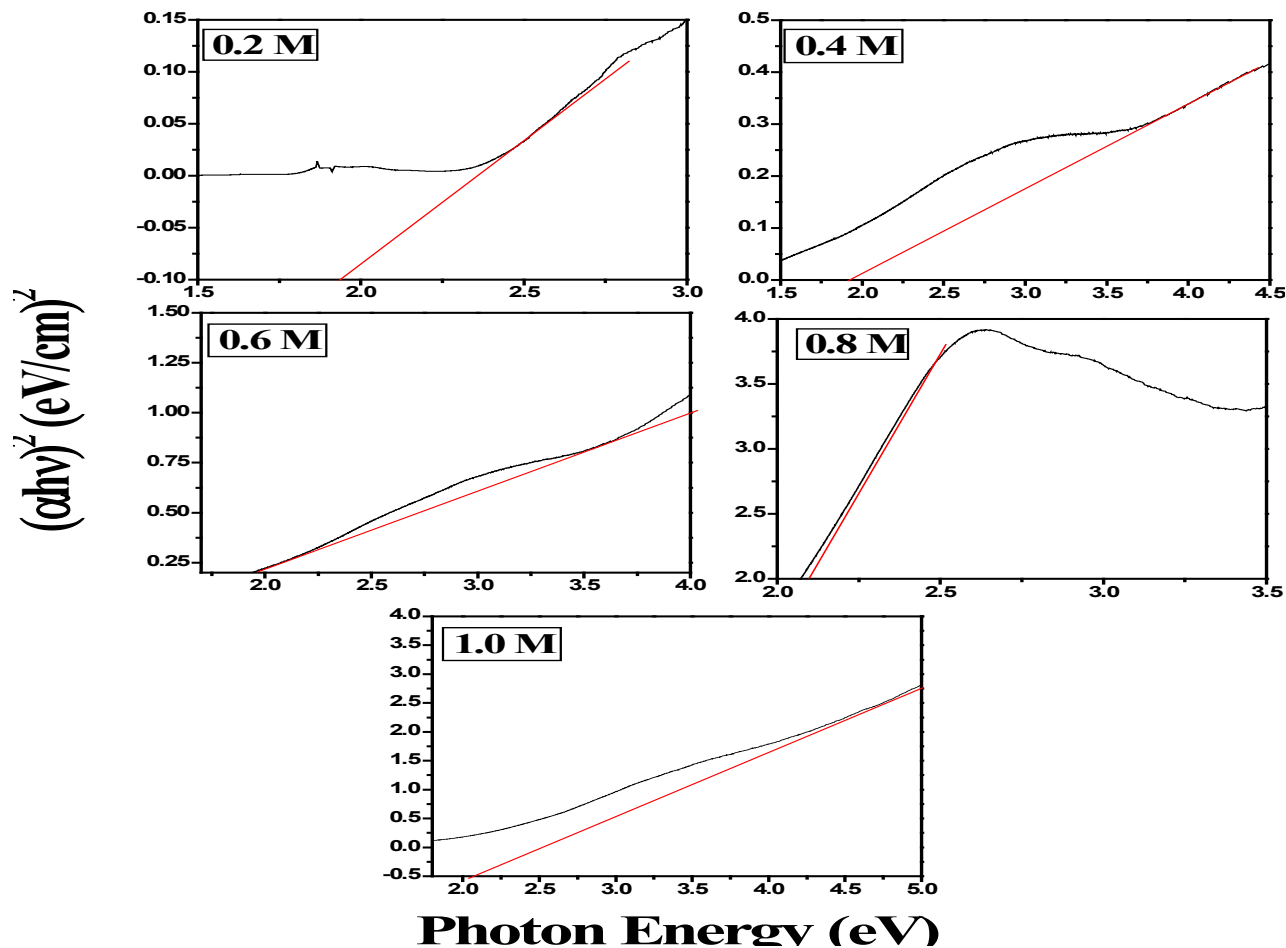


Figure 6. Estimation of optical band gap of akaganeite nanoparticles.

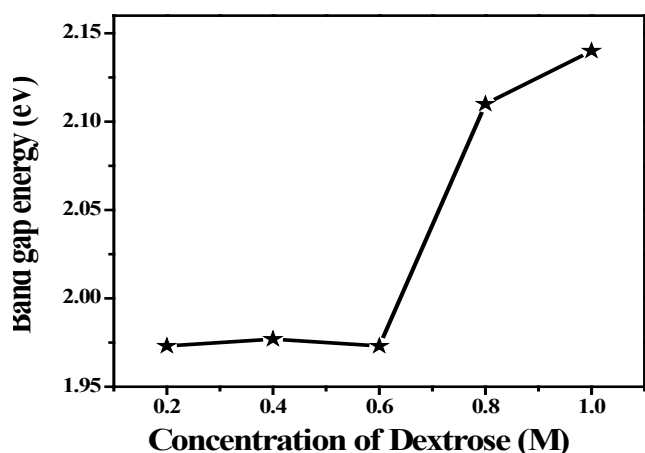


Figure 7. Variation of band gap energy versus concentration of dextrose.

The bandgap energy of akaganeite nanoparticles is found to vary in the range of 1.97 to 2.14 eV which is in agreement with the earlier reports[22,46,47]. Thus the akaganeite nanoparticles can become a useful member of a wide bandgap semiconductor family. The estimated E_g value is plotted with respect to the concentration of dextrose as shown in Figure. 7. It is seen from the Figure. 7 that the bandgap increases steeply for 0.6, 0.8 and 1.0 M of dextrose. A possible correlation is that the particles have undergone size reduction for these molar ratios only and they are pure akaganeite nanoparticles free from other neighbouring phases. It is established that the size reduction of nanoparticles of akaganeite is responsible for optical bandgap widening which is a widely accepted concept of wide bandgap semiconductor nanoparticles[46-49].

3.4. VSM Study

The hysteresis curves of akaganeite nanoparticles prepared with different molar ratios of dextrose are shown in Figure.8. It is seen from Figure.8 that the loops obtained for 0.2 and 0.4 M, show saturation after remanence magnetization at a maximum applied field of 2 T.

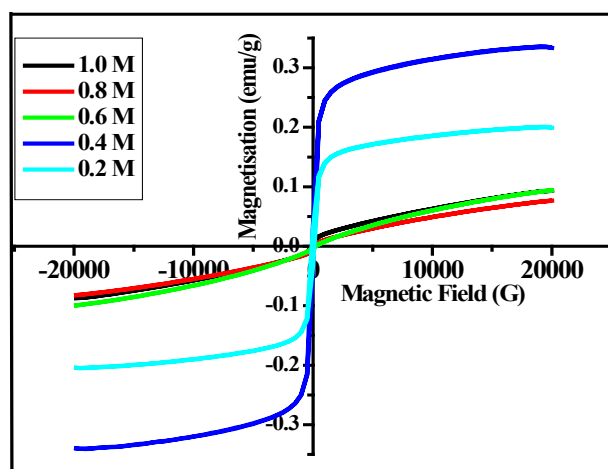


Figure 8. Hysteresis loops of akaganeite nanoparticles.

For molar ratios of dextrose 0.6, 0.8, and 1.0 M, the magnetization of the particles does not saturate at the maximum limit of field. However the bend in the nature of the loops

for these three cases clearly indicates that the nanoparticles are not in paramagnetic state. It is also quite contradictory to the earlier report[50] of the antiferromagnetic nature of akaganeite only below 200 K. The magnetic parameters estimated from the hysteresis loops are given in Table 3. Moreover the observed values of coercivity and magnetization in Table 3 are quite appreciable for the pure akaganeite nanoparticles though smaller than those of mixed phases. An ordered magnetic phase is expected to arise due to the arrangement of Fe^{3+} ions in the lattices of nanoparticles. For nanosize of the particles, a substantially higher number of magnetic moment bearing ions exist on their surfaces than the cores. It is explained that a small resultant moment is caused by uncompensation of them[51]. Therefore it is explained that the observation of hysteresis loop in the present akaganeite particle is purely due to the outcome of the difference in the alignment of magnetic moments in nano-regime i.e., an explicit display of core shell effect. Thus we furnish the first ever report on ferrimagnetic akaganeite nanoparticles at room temperature. Further the particle size as reported by XRD data is approaching the dimension equivalent to that of superparamagnetic phase. The observed low values of magnetization and coercivity could be attributed to the coexistence of ferrimagnetic phase with an admixture of superparamagnetic domains in akaganeite nanoparticles.

Table 3. Magnetic properties of akaganeite nanoparticles.

Concentration of Dextrose M	Magnetization M_s (emu/g)	Coercivity H_c (G)
0.2	5.63	98.53
0.4	9.95	62.52
0.6	2.85	22.14
0.8	2.65	46.97
1.0	1.92	32.26

3.5. EPR Study

EPR spectra of akaganeite nanoparticles are shown in Figure 9. It is seen from Figure.9 that the spectra of 0.2 and 0.4 M show single and symmetric signal whereas those of 0.6, 0.8 and 1.0 M are found to possess two resonance bands. A band with higher intensity appears at higher field and other one of lower intensity is found at lower fields. The g values estimated from the resonance bands and given in Table 4 show two distinctly different sets. The g -values of high field bands are observed to be more than 2 whereas the low field bands have g values close to 2. The origin of the peaks observed at lower field for the 0.6, 0.8 and 1.0 M samples is clearly due to spins of Fe^{3+} ions in akaganeite[52]. Another EPR signal at higher field is explained to possess distortions caused by the environment on Fe^{3+} ions since no other EPR ion is present in akaganeite structure.

It is observed that the peak to peak line width (ΔH_{pp}) increases with the increase in dextrose concentration. It may be recalled that the increase in dextrose concentration decreases the particle size. Therefore the coupling force inside particles is expected to play major role in affecting the line width of the samples. It is also observed that the resonance

bands are shifted to lower fields with decrease in particle size. In the case of nanoparticles, the peak to peak line width and the resonance field are often correlated. With the decrease in particle size the time of spin fluctuations increases which results in a continuous and simultaneous line broadening and resonance shift towards lower resonance fields.

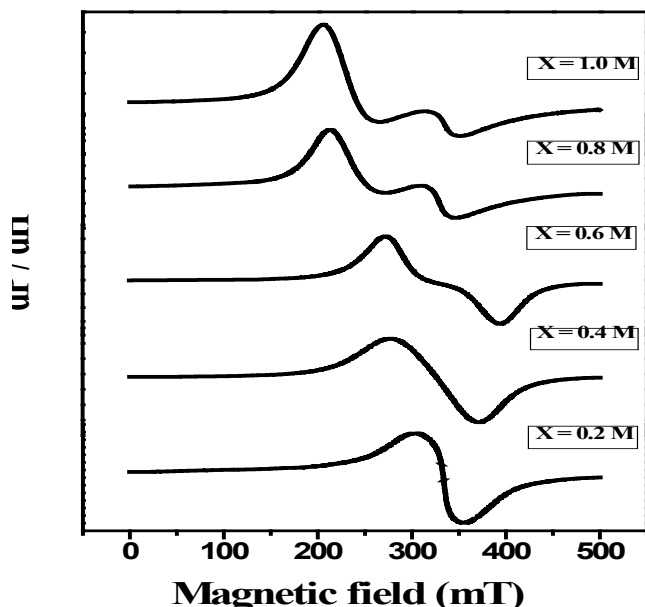


Figure 9. EPR spectra of akaganeite nanoparticles.

Table 4. EPR parameters of akaganeite nanoparticles.

Concentration of Dextrose (M)	g- factor		ΔH_{pp} (G)		Relaxation time in sec	
	g_1	g_2	ΔH_{pp1}	ΔH_{pp2}	$\tau_1 \times 10^{-10}$	$\tau_2 \times 10^{-10}$
0.2	2.004	-	879.37	-	4.057	-
0.4	2.082	-	1653.76	-	2.076	-
0.6	2.289	1.822	1199.14	879.37	2.603	4.464
0.8	2.913	2.014	1065.80	612.80	2.302	5.791
1.0	2.843	2.050	983.49	655.60	2.557	5.310

The spin-spin relaxation process accounts for the energy difference transferred to the neighboring electrons and it plays an important role in limiting the line width. The spin-spin relaxation time with respect to different concentration of dextrose is given in Table 4. It is seen that the behavior of relaxation time is converse to that of line width. For microwave applications long relaxation time is needed [53]. The order of relaxation time observed for the present akaganeite nanoparticles informs that they can serve as one of the potential candidates for the microwave applications.

4. Conclusions

The synthesis of akaganeite (β -FeOOH) nanoparticles using dextrose as chelating agent has been successfully investigated with a focus on optimization of a lone phase, free from secondary phases, and stabilize it. The XRD patterns confirm the setting in of single phase akaganeite for the

concentration of dextrose above $X = 0.6M$ although secondary phase co-existed in the low concentrations for $X = 0.2$ and $0.4M$. It helps to conclude that the sufficient amount of dextrose during the precipitation process has ensured the formation of pure akaganeite nanoparticles. DSC study reports that more quantity of dextrose in the solution resolves the akaganeite phase from other immediate phases. FTIR spectra of akaganeite nanoparticles bring out a decrease in bond length with increase in the concentration of dextrose during the preparation which enlightens the role of dextrose in speeding up the growth mechanism of nanoparticles. UV-Visible spectroscopic study reports that the bandgap energy of akaganeite lies between 1.97 and 2.14eV and hence adds a new member in the list of wide bandgap semiconductors. The superparamagnetic domains present in the akaganeite nanoparticles would have a real value for pharmaceutical and biomedical applications like targeted drug delivery. The large values of relaxation time and peak to peak line width observed in EPR study assure a future scope for these akaganeite nanoparticles – a wide band gap semiconductor- in microwave region.

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