

# Adsorption of Cadmium Ions on Geopolymers Derived from Ordinary Clay and Rice Husk Ash

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**Abstract** The presence of heavy metals in water is of growing concern. One crucial global strain is the existing disparity in water distribution resulting into millions of people lacking access to safe clean water. This is exasperated by the looming climate change and increased pollution impendence posed by industrialization and population growth. Consequently, pursuance of sustainable materials and technologies for wastewater remediation is critical. Adoption of adsorbents for water treatment is desirable due to their simplicity and cheap materials involved. Geopolymers are prepared hydrothermally using alumina and silica gleaned from clay and rice husks waste materials respectively. These materials are locally attainable and environmentally benign. This work reports synthesis of geopolymers GP-1, GP-2 and GP-3 prepared using clays from three different regions. The XRD diffraction patterns obtained revealed the presence of quartz ( $\text{SiO}_2$ ), albite ( $\text{NaAlSi}_3\text{O}_8$ ), vermiculite ( $\text{Mg,Fe,Al}_3((\text{Al,Si})_4\text{O}_{10})(\text{OH})_2 \cdot 4\text{H}_2\text{O}$ ), microcline ( $\text{KAlSi}_3\text{O}_8$ ), analcime ( $\text{NaAlSi}_2\text{O}_6 \cdot \text{H}_2\text{O}$ ) and natrite ( $\text{Na}_2\text{CO}_3$ ) among other crystalline phases. The initial adsorption rates were found to increase as the Si/Al ratio of the geopolymers increased. The adsorption rates were found to fit into the second order kinetic model. This implies that Cd (II) ions adsorb on geopolymer adsorbents through chemisorption.

**Keywords** Geopolymers, Characterized, Equilibrium, Thermodynamic, Kinetics

## 1. Introduction

Rapid industrialization and technological advancement experienced in Kenya has led to increased use of cadmium metal. This metal is extremely toxic and of significant environmental and occupational concern [1] due to lack of homeostatic control in human body. This metal is one of the main causes of various human pathologies such as kidney damage, liver disorder, high blood pressure, stomach problems among other ailments [2]. Geopolymers have gained interest for use as adsorbents in wastewater treatment. They can be easily synthesized from simple clays, aluminosilicate and rice husk ash [3]. They consists of polymeric Si–O–Al and Si–O–Si frameworks joined together by sharing oxygen atoms to form a three dimensional network [4]. Moreover, in the approach to replace the traditional adsorbents, geopolymers are desirable because they are green materials that are easily prepared by a simple chemical reaction of readily available raw source of aluminosilicate [5]. For this reason, geopolymers have started being used as adsorbents within the procedures of elimination of metal ions from the wastewater as an alternative to commercial adsorbents within the industrial

sector [6]. Preceding studies showed that higher Si: Al ratio caused increment in silica content within the framework that caused increased negatively charged groups of O–Si–O– within the geopolymer [7]. This study reports synthesis, characterization, their kinetics and thermodynamic effects of geopolymers on their use in the adsorption process of cadmium ions from aqueous phase.

## 2. Experimental Procedures

### 2.1. Chemicals and Reagents

NaOH was used as an alkaline catalyst during geopolymer synthesis.  $\text{Cd}(\text{NO}_3)_2$  and  $\text{HNO}_3$  were used in adsorption experiments and were of analytical grade supplied by Sigma Aldrich.

### 2.2. Clay and Rice Husk Materials

The clays used in this study were acquired from Kakamega, Kuresoi and Molo in Kenya. They were transported to Kenyatta university laboratory, washed with distilled water, then air dried for 2 weeks. Calcination was done at  $700^\circ\text{C}$  in a furnace (Eisklo LN120) and later ground to powder with the use of grinder (Reutsch SR 200). Calcined clay was used as a supplier of  $\text{Al}_2\text{O}_3$ . Rice husk acquired from a rice processor in Mwea location in Kenya was washed with distilled water for 20 min at room temperature, using a liquid/solid ratio of 15 L: 1 kg. The

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solids were allowed to settle and dried at 60°C in an oven. Leaching was then achieved with 3% HCl solution, using reflux for 2 hours. The solids, after separation by means of settling and filtration, were washed with distilled water, to do away with the acid retained, and then dried at 60°C. Calcination was accomplished at 600°C for 2 hours using a muffle furnace to obtain ash which was used as a source of silica [8].

### 2.3. Preparation of Geopolymers

Geopolymers were prepared following the procedure given by López *et al* [7] with slight modifications. Rice husk ash was blended with 8M NaOH solution in the liquid/solid ratio of 5 L: 1 Kg and stirred with the aid of a magnetic stirrer for 15 minutes to form an alkaline activator. According to Mustafa Al Bakri *et al.* [9], these reactions occur at a high rate when the alkaline activator contains soluble silicates. 50 g of clay powder was added to the mixture and thoroughly mixed using a ball mill for a period of 5 hours. The resultant pastes were vibrated for 5 minutes in pestles. The resulting pastes were then cured at 80°C for 12 hours in order to start the condensation reaction. The pastes were then removed from the molds, placed in an oven at 200°C for 12 hours in order for the polycondensation process to complete [7].

### 2.4. Characterization of Geopolymers

Characterization of prepared geopolymer was done before the adsorption experiments using FT-IR (IRTracer 100 Shimadzu) for determination of molecular vibrations of the functional groups in the geopolymers. The crystal structures (mineral phases present) were determined using X-ray diffractometer (XRD) (D2 Phaser Bruker), X-ray fluorescence (XRF) (S1 Titan) to determine the chemical composition of principal components while SEM (Zeiss Evo LS 15) was used for microstructure analysis [7].

### 2.5. Adsorption Experiments

#### 2.5.1. Effect of pH on Adsorption of Cd (II) ions

Experiments were administered by adjusting pH from 2 to 6 of Cd (II) ions solution while maintaining all other experimental conditions constant. The pH modification was done using procedures given by Moreno-piraján and Giraldo [10]. 50 mL pH accustomed containing 100 mg/L of Cd (II) ions in monocomponent system was placed in 100 mL stoppered plastic bottles. 0.10g of the geopolymer materials were accurately weighed and put into stoppered bottles. The bottles were then shaken continuously for one hour placed in a thermostatic electric shaker (SHR-2D) at 120 rpm which was maintained at 25°C ± 1. All the experiments were carried out in triplicates and mean average values used for further calculations. The reagent bottles were then successively withdrawn and the supernatant solutions separated by filtering through whatman no 41 filter papers. The resultant solution was analyzed for cadmium ions using

a FAAS (AA 6300 Shimadzu) following the procedure reported by Ryan *et al.* [11].

#### 2.5.2. Kinetic Studies of Adsorption Process

For the batch kinetic experiments, about 0.10 g of the geopolymer adsorbents were placed in 100 mL stoppered bottles. 50 mL of 100 mg/L cadmium ion solutions were added to each bottle and the mixture placed in a water bath electric shaker at residence time ranging from 20 to 80 minutes. The mixtures were subjected to a constant shaking speed of 120 rpm at 25°C ± 1 and then filtered. The metal ions concentration contained in the filtrate were measured using AAS spectrophotometer. Plots of log (q<sub>e</sub> - q<sub>t</sub>) against t and t/q<sub>t</sub> against t were used to determine kinetic parameters for Pseudo- first order and Pseudo- second order as per equations 1 and 2 respectively.

$$\text{Log } (q_e - q_t) = \text{Log } q_e - \frac{K t}{2.303} \quad (1)$$

$$\frac{t}{q_t} = \frac{1}{K_2 q_e^2} + \frac{1}{q_t} t \quad (2)$$

Where q<sub>e</sub> is amount of cadmium ions adsorbed at equilibrium while q<sub>t</sub> is amount adsorbed at a given time, t is residence time in minutes; K and K<sub>2</sub> are the rate constants for Lagergren and Pseudo- second order respectively. The data obtained from these graphs was used to plot the intraparticle diffusion models using equation 3.

$$R = K_{id} (t)^a \quad (3)$$

Where R = percent of Cd (II) adsorbed, t = contact time (minutes), a = gradient of linear plot [12].

#### 2.5.3. Thermodynamic Studies of Adsorption Process

Batch sorption experiments at different temperatures were done to investigate the effect of temperature on thermodynamic parameters. Stoppered bottles containing 0.10 g of geopolymer adsorbents and 50 mL of 100 mg/L cadmium aqueous solutions of pH 5.0 were used. The stoppered bottles and their contents were placed on a water bath electric shaker at various temperatures of 293, 298, 308, 318 and 328 K and shaken for one hour. The adsorbent/adsorbate mixture was then filtered and final concentration of metal ions in the filtrates analyzed using AAS. The data obtained was used to tabulate the thermodynamic parameters such as the adsorption standard free energy changes (ΔG°), the standard enthalpy change (ΔH°) and the standard entropy change (ΔS°) following equations 4 and 5 [13].

$$\Delta G^\circ = -RT \ln K \quad (4)$$

$$\ln K = \Delta S^\circ / R - \Delta H^\circ / RT \quad (5)$$

where R (8.3145 J mol<sup>-1</sup> K<sup>-1</sup>) is the gas constant, K (L g<sup>-1</sup>) is the Langmuir constant and T (K) is the absolute temperature [13].

### 3. Results and Discussion

#### 3.1. Chemical Composition of Geopolymer

The components of the geopolymer adsorbents GP-1, GP-2 and GP-3 were mainly silica and alumina as shown in table 1. Silica, SiO<sub>2</sub> constituted 66.84, 66.21 and 73.99% while alumina, Al<sub>2</sub>O<sub>3</sub> constituted 16.05, 15.83 and 8.93% of GP-1, GP-2 and GP-3 respectively. Additional elements found in lesser quantities include Fe at 6.42, 5.65 and 9.05%, K<sub>2</sub>O at 4.43, 6.08 and 2.89%, CaO at 1.88%, 1.43% and 0.92% and Ti at 1.38, 1.12% and 1.85% for GP-1, GP-2 and GP-3 respectively. The composition is in agreement with results obtained by Lopez *et al.* [7]

**Table 1.** Percentage chemical composition of geopolymers

Composition (%)	GP-1	GP-2	GP-3
	Mean% ± SD	Mean% ± SD	Mean% ± SD
SiO <sub>2</sub>	66.84 ± 0.44	66.21 ± 0.49	73.99 ± 0.45
Al <sub>2</sub> O <sub>3</sub>	16.05 ± 0.39	15.83 ± 0.43	8.93 ± 0.32A
Fe	6.42 ± 0.03	5.65 ± 0.03	9.05 ± 0.04
CaO	1.88 ± 0.02	1.43 ± 0.02	0.92 ± 0.02
MgO	-	1.21 ± 1.52	-
Mn	0.11 ± 0.01	0.09 ± 0.01	0.07 ± 0.01
K <sub>2</sub> O	4.43 ± 0.03	6.08 ± 0.04	2.89 ± 0.02
P <sub>2</sub> O <sub>5</sub>	-	0.64 ± 0.05	0.61 ± 0.05
Ti	1.38 ± 0.01	1.12 ± 0.01	1.85 ± 0.01

GP-1= Geopolymer from Kakamega clay, GP- 2= Geopolymer from Kuresoi clay, GP- 3= Geopolymer from Molo clay and SD = Standard deviation.

#### 3.2. FTIR Spectra Analysis on Geopolymers

The infra-red spectra of the various geopolymer adsorbents GP-1, GP-2, and GP-3 are shown in Figure 1. The broad bands at ~3400 cm<sup>-1</sup> are attributed to -OH stretching mode. The spectral band at around 1410 cm<sup>-1</sup> is characteristic of the asymmetric O-C-O bonds of CO<sub>3</sub><sup>2-</sup> stretching mode. This is associated with sodium bicarbonate formed due to atmospheric carbonation of alkaline NaOH aqueous phase and diffused on the geopolymeric materials

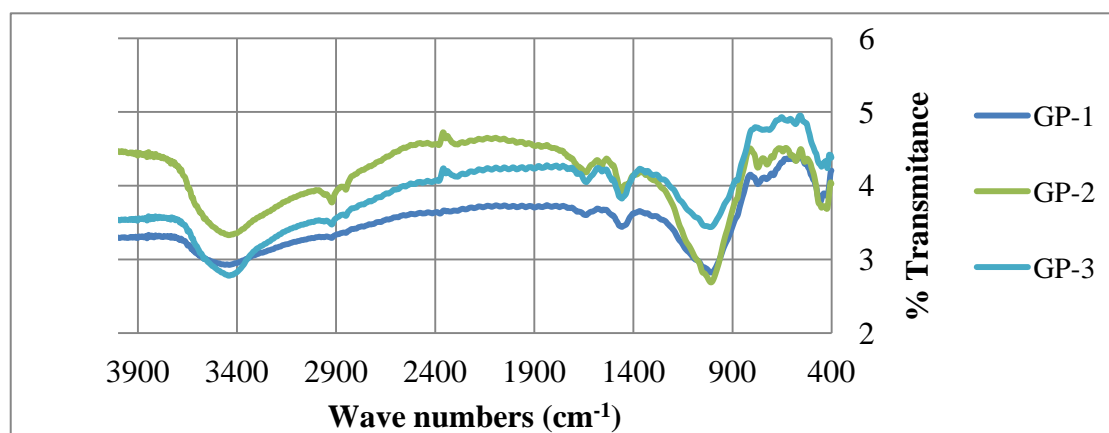
surface [14, 15]. The characteristic Al-O and Si-O asymmetric stretching vibrations appear as strong peak at ~979 cm<sup>-1</sup> [16]. The band in the region of 775-650 cm<sup>-1</sup> attributed to symmetrical vibrations of tetrahedral groups (TO<sub>4</sub>) and at ~460 cm<sup>-1</sup> is assigned to in-plane bending of Al-O and Si-O linkages [16].

#### 3.3. SEM Microstructure Analysis of Geopolymers

Figure 2 shows the SEM micrographs of GP-1, GP-2, GP-3 and EDS. The images reveal that the geopolymeric adsorbents have compacted irregularly shapes. The cavities found in micrograph for GP-2 indicate that it consists of tubular vitreous network [17]. EDXRF analysis of gel showed that gel mostly consists of the phases containing Na-Si-Al in the bulk region. This is suggestive of formation of silicate-activated gel all over the inter particles volume [3, 18]. The major elements (e.g., Na, Al, Si, and O) make up geopolymers; Ca, Mg and Fe are also present and have influence in geopolymerization process [19]. These remnants (Fe, Ca, K, Mg) obviously represent the calcined clay and rice husk ash phases, which for various reasons, did not dissolve during alkali activation [17]. Literature report by Lloyd *et al.* [18]; suggested that during alkaline activation the remnants may disperse through the gel. This correlates with the published works of Lee and Van Deventer [14].

#### 3.4. XRD Characterization of Geopolymers

Quartz (SiO<sub>2</sub>), albite (NaAlSi<sub>3</sub>O<sub>8</sub>), vermiculite (Mg, Fe, Al)<sub>3</sub>((Al,Si)<sub>4</sub>O<sub>10</sub>)(OH)<sub>2</sub>·4H<sub>2</sub>O, microcline (KAlSi<sub>3</sub>O<sub>8</sub>), analcime (NaAlSi<sub>2</sub>O<sub>6</sub>·H<sub>2</sub>O) and natrite (Na<sub>2</sub>CO<sub>3</sub>) were the major mineral phases identified by XRD diffraction patterns as shown in figure 3. The formation of natrite (Na<sub>2</sub>CO<sub>3</sub>) is attributed atmospheric CO<sub>2</sub> reacting with the sodium-rich pore solutions [20]. The characteristic amorphous geopolymer broad hump peak between 18-36° (2θ) [21, 22] was found in all the XRD patterns but was more pronounced in GP-3.



**Figure 1.** Shows the FT-IR spectra for geopolymer materials GP-1, GP-2, and GP-3

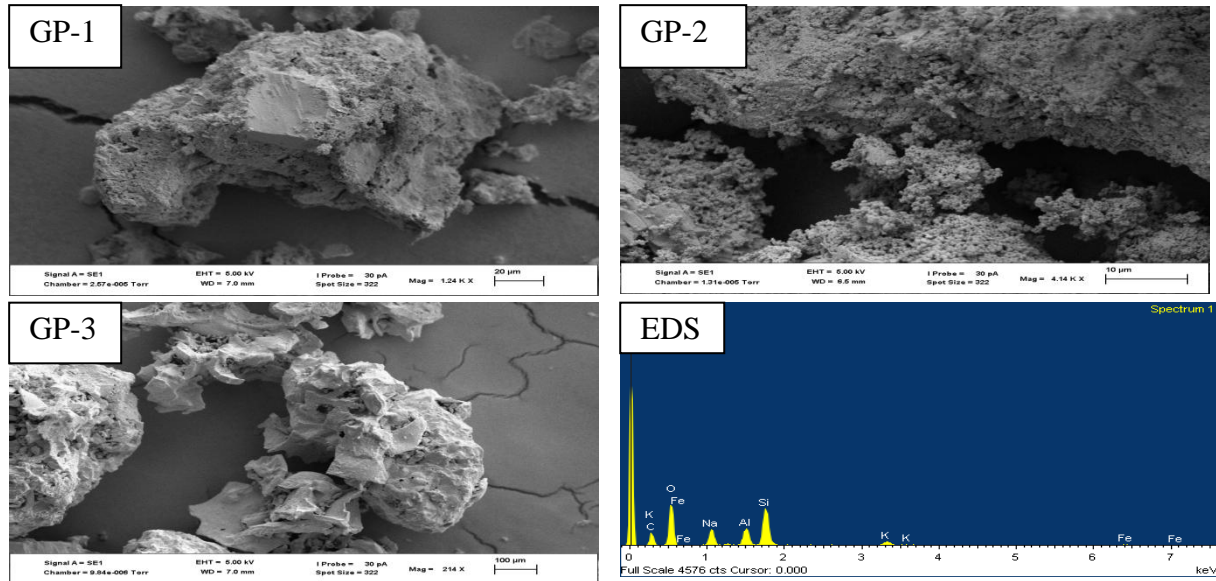


Figure 2. Shows the SEM micrographs for GP-1, GP-2, GP-3 and EDS spectrum

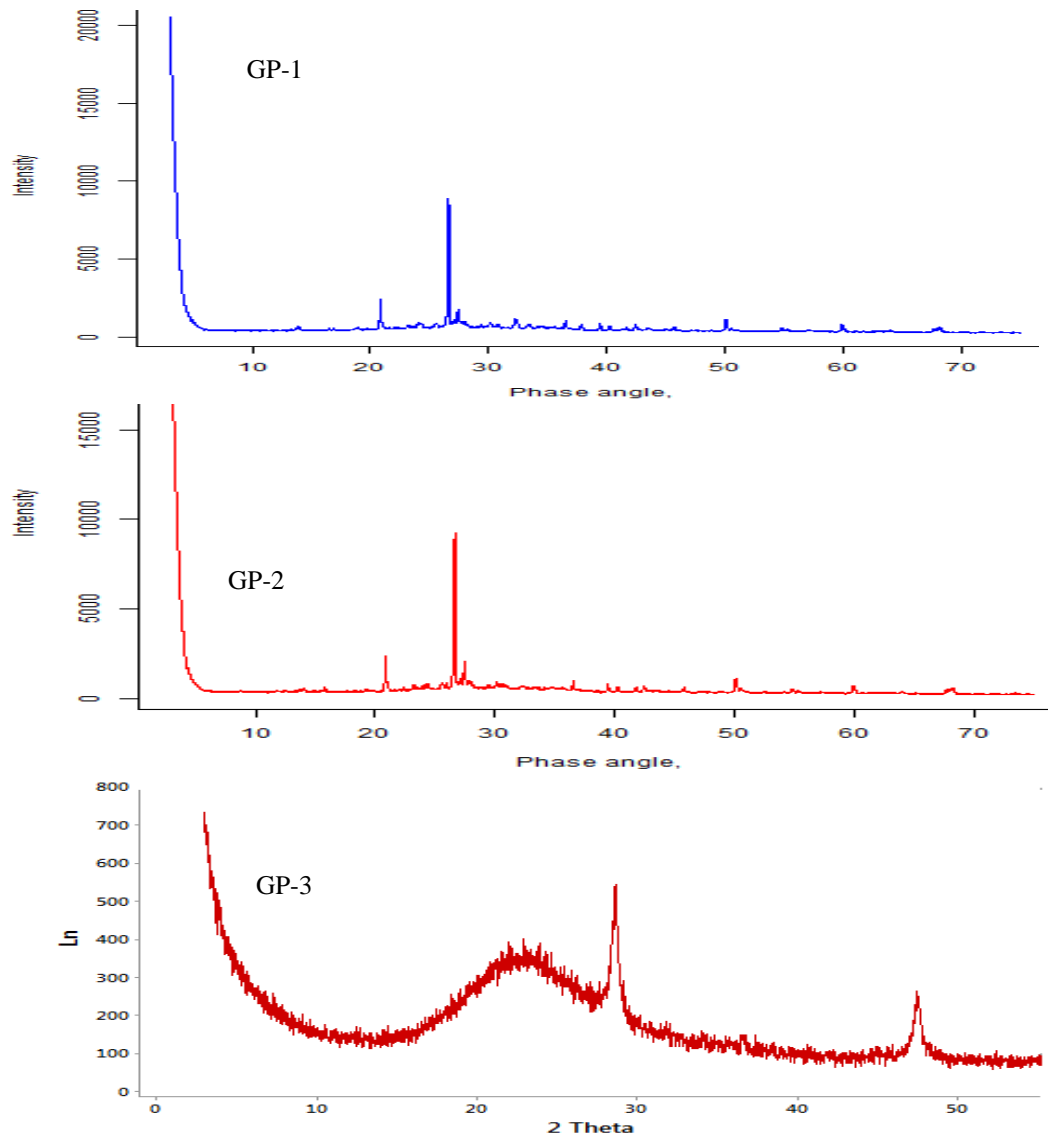


Figure 3. Shows diffraction patterns for geopolymeric materials GP-1, GP-2 and GP-3

### 3.5. Adsorption Studies of Geopolymers

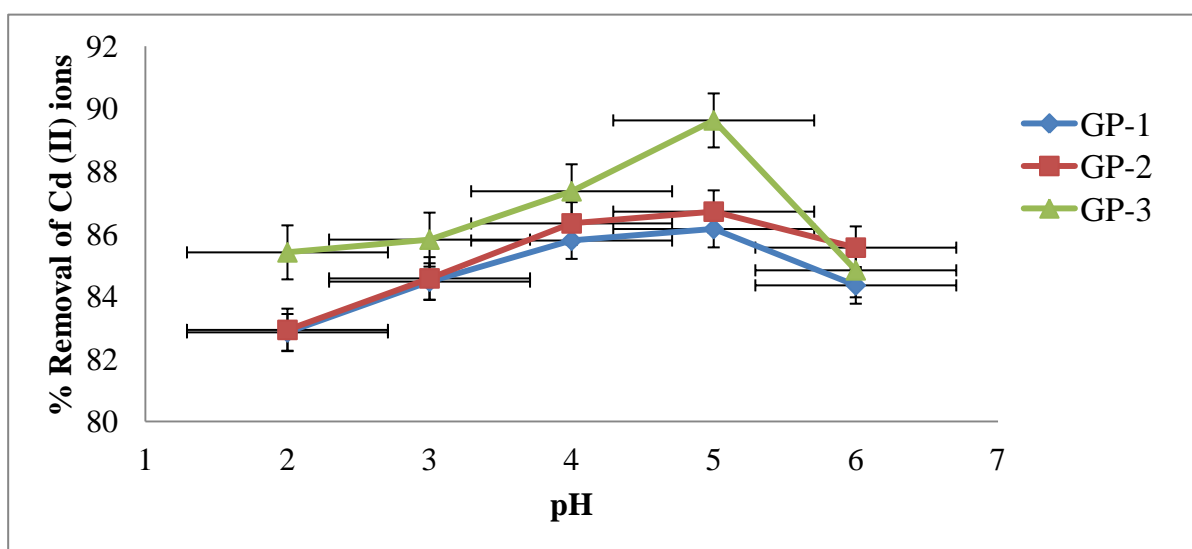
#### 3.5.1. Effect of pH on Adsorption of Cd (II) ions

The adsorption efficiency of the geopolymeric adsorbents for Cd (II) ions at various pH conditions are expressed in figure 4. The experiments were done using 0.1 g of adsorbent and 50 mL Cd (II) ions solutions at concentrations of 100 mg/L while the pH was varied from 2 to 6. The mean percentage removal increased from  $82.85\% \pm 0.20$ ,  $82.93\% \pm 0.13$  and  $85.41\% \pm 0.11$  to  $86.16\% \pm 0.31$ ,  $86.71\% \pm 0.26$  and  $89.63\% \pm 0.18$ , when pH was varied from (pH2–5), for geopolymer materials GP-1, GP-2 and GP-3 respectively. The data obtained revealed that optimum adsorption was achieved at pH 5. Low adsorption at low pH was attributed to competition caused by hydronium ions on the sorption sites.

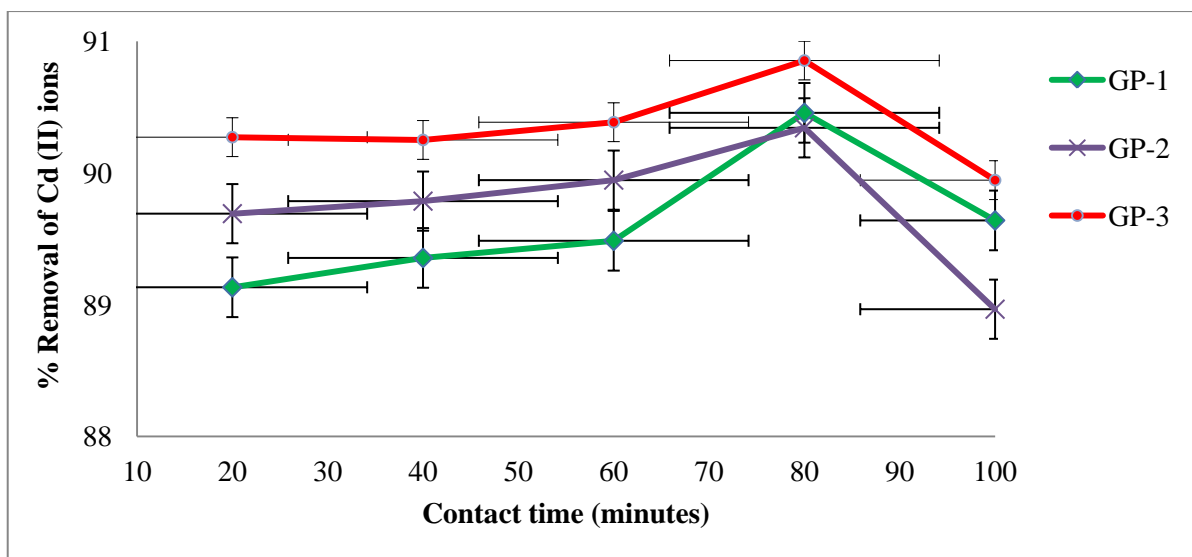
This concurs with similar work on adsorption of cadmium reported by Kounou *et al.* [23]. Adsorbent GP- 3 showed higher efficiency for cadmium uptake than GP-1 and GP-2 at all pH values. Subsequently, pH 5.0 was chosen as the working value for all the other batch experiments.

#### 3.5.2. Effect of Contact Time

The adsorption of Cd (II) ions by the geopolymeric adsorbents increases with increased to an optimum mean percentage uptake of  $90.46\% \pm 0.48$ ,  $90.34\% \pm 0.12$  and  $90.85\% \pm 0.06$  for GP-1, GP-2 and GP-3 respectively as shown in figure 5. The extent of sorption increased expeditiously till the optimum equilibrium values were reached at 80 min for all the geopolymers in concurrence with literature report by Zhang *et al.* [24].



**Figure 4.** Effect of pH on Cd (II) removal onto geopolymer materials (Cd (II) concentration=100 mg/L, adsorbent dose= 0.1 g/50 mL, shaking speed = 120 rpm, temperature = 25 °C and error bars indicate the standard error from the mean)



**Figure 5.** Effect of contact time on Cd (II) removal onto geopolymer materials (Cd (II) concentration=100 mg/L, adsorbent dose= 0.1 g/50 mL, shaking speed = 120 rpm, pH = 5.0 temperature = 25°C and error bars with standard error from the mean)

### 3.5.3. Effect of Temperature

Figure 6 represents the change in extent of adsorption of cadmium with respect to temperature. The experiments were done by varying the temperature from 293 to 328K while keeping the rest of the parameters such as adsorbate concentration at 100 mg/ L, the pH at 5, and the adsorbent concentration at 0.1 g in 50 mL. The mean percent removal increased from  $86.83\% \pm 0.11$ ,  $89.22\% \pm 0.19$  and  $90.49\% \pm 0.03$  to  $90.61\% \pm 0.07$ ,  $91.52\% \pm 0.48$  and  $95.04\% \pm 0.02$  when temperature was adjusted from 293 to 328 K for adsorbents GP-1, GP-2 and GP-3 respectively. The increment of adsorption with temperature is attributed to increased ions mobility, thus increasing the number of ions that interact with active sites of the adsorbent surfaces [25]. Similar trends have been reported by other researchers for aqueous phase adsorption [26].

### 3.5.4. Adsorption Thermodynamics for Cd (II) ions onto Geopolymeric Adsorbents

The thermodynamic parameters for the adsorption of Cd (II) ions onto the geopolymeric adsorbents at the various temperatures are summarized in table 2. The values for changes in free energy,  $\Delta G^\circ$  were negative for all the adsorbents at all temperatures indicating spontaneity of process [27]. The values became more negative ranging from  $-4.60 \pm 0.02$ ,  $-5.49 \pm 0.01$  and  $-5.15 \pm 0.05$  at 293 K to  $-6.00 \pm 0.02$ ,  $-7.81 \pm 0.01$  and  $-6.06 \pm 0.37$  at 318 K for GP-1, GP-2 and GP-3 respectively. This is expected as rise in temperature increases the spontaneity of adsorption [28]. The decreased values of  $\Delta G$  with increase temperature imply that the reaction is more spontaneous and favourable at higher temperature [29]. The values for entropy change were all positive and significant signifying that enormous structural changes took place at the active sites of the geopolymers during adsorption process. This is expected because for metal ions to be adsorbed, they have to get rid of part of their hydration sheath. The endothermicity of the dehydration energy supersedes the corresponding exothermicity of the ions getting attached to the surface [30].

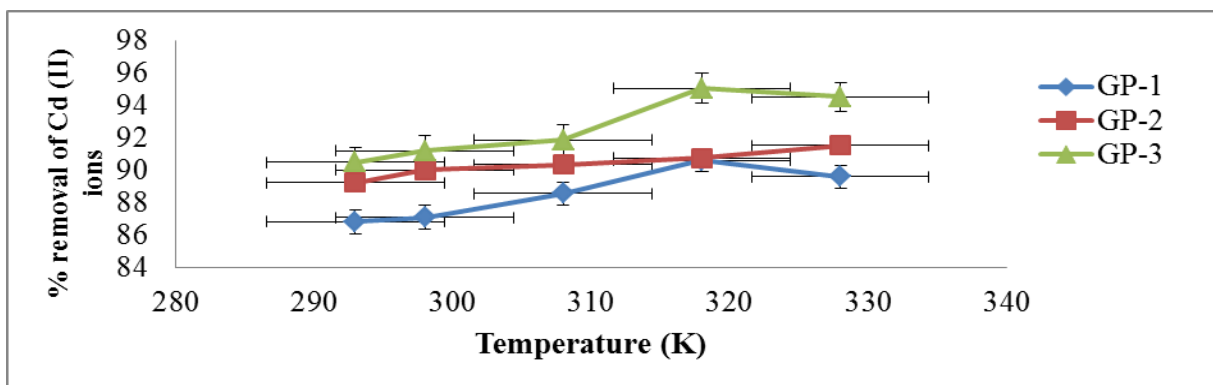
### 3.5.5. Kinetic Modeling for Adsorption of Cd (II) ions

The pseudo-first and pseudo-second order kinetic parameters calculated by linear regression for Cd (II) ions are summarized in table 3. As indicated in the table, the data correlated closely with pseudo- second order kinetics model with  $R^2$  coefficient values of 1.0, 1.0 and 0.999 for adsorbents GP-1, GP-2 and GP-3 respectively. The initial adsorption rate ( $h_0$ ) obtained were 42.37 to 54.07 and 72.37 mg /g min as the  $\text{SiO}_2/\text{Al}_2\text{O}_3$  ratio of geopolymer increased from 4.16 to 8.28 for adsorbents GP-1, GP-2 and GP-3 respectively. This is due to the increased driving force of Cd (II) ions from the liquid to the solid phase [31].

The first order rate constant  $k_1$  and equilibrium adsorption density  $q_e$  were determined from the slopes and intercepts of plots of  $\log (q_e - q_t)$  versus  $t$  and are shown in table 3. As indicated in the table, the calculated theoretical data deviated considerably from the experimental data. Further, the theoretical  $q_e$  values found from the pseudo-first order model did not give reasonable values. The correlation coefficients obtained were greater than 0.99. The calculated  $q_e$  values for the second order kinetic model agreed well with the experimental values. Thus, the adsorption system studied belonged to the pseudo-second order model. This implies that Cd (II) ions adsorption on geopolymer adsorbents occurs through chemical processes involving the valence forces or the shared exchangeable electrons [32]. This means that the chemisorption reaction or activation process is more predominant in the rate-controlling step for the cadmium system. Similar phenomena have been reported in the adsorption of copper and lead on tea waste [33].

### 3.5.6. Intraparticle Diffusion Models of Cd (II) Uptake onto Geopolymers

The intraparticle plots of Cd (II) ions data exhibited multi linear plots as shown in figure 7. This implies that the adsorption process is governed by two or more steps. The first linear part is likely attributed to the abrupt exploitation of the readily available sorbing sites on the sorbent surface while the second portion possibly credited to the very slow diffusion of Cd (II) ions from the geopolymer surface sites into the inner pores [34].



**Figure 6.** Effect of temperature on Cd (II) removal onto geopolymer materials (Contact time = 60 minutes, shaking speed of 120 rpm, pH = 5.0 and adsorbent dose = 0.1g/50mL)

**Table 2.** Adsorption thermodynamic parameters of Cd (II) ions

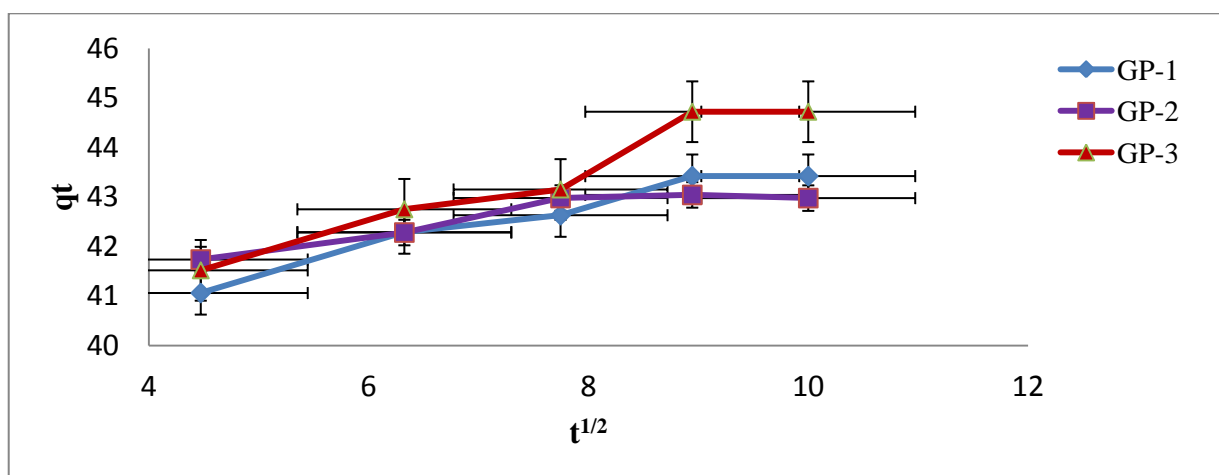
GP	$\Delta H^\circ$ (KJ/Mole)	$\Delta S^\circ$ (J/Mol/K)	$\Delta G^\circ$ (kJ/Mole)			
			293K Mean $\pm$ SD	298K Mean $\pm$ SD	308K Mean $\pm$ SD	318K Mean $\pm$ SD
1	7.96	42.87	-4.60 $\pm$ 0.02 <sup>d</sup>	-4.73 $\pm$ 0.01 <sup>c</sup>	-5.24 $\pm$ 0.05 <sup>b</sup>	-6.00 $\pm$ 0.02 <sup>a</sup>
2	16.07	73.43	-5.49 $\pm$ 0.01 <sup>d</sup>	-5.80 $\pm$ 0.01 <sup>c</sup>	-6.21 $\pm$ 0.08 <sup>b</sup>	-7.81 $\pm$ 0.01 <sup>a</sup>
3	5.37	36.04	-5.15 $\pm$ 0.05 <sup>c</sup>	-5.45 $\pm$ 0.01 <sup>bc</sup>	-5.73 $\pm$ 0.09 <sup>ab</sup>	-6.06 $\pm$ 0.37 <sup>a</sup>

Means in a row with the same letter are not significantly different from each other (Tukey–Kramer test,  $P > 0.05$ ), SD= standard deviation, GP= geopolymer,  $\Delta H^\circ$  = Standard enthalpy change,  $\Delta S^\circ$  = Standard entropy change and  $\Delta G^\circ$  = Standard free energy change.

**Table 3.** Descriptive data on rate constants of Cd (II) adsorption onto geopolymers as estimated from Pseudo- first and second order kinetic models

GP	Pseudo-first order						Pseudo-second order			
	$q_e(\text{exp})$	SD	$q_e(\text{cal})$	SD	$K_1$	$R^2$	$q_e(\text{cal})$	SD	$h_0$	$R^2$
1	43.23	0.07	2.46	0.53	0.013	0.729	43.48	0.20	42.37	1.0
2	43.14	0.05	1.79	0.27	0.025	0.509	43.23	0.11	54.07	1.0
3	44.73	0.04	4.65	0.14	0.019	0.991	45.18	0.12	72.37	0.999

GP= Geopolymer adsorbents, SD= Standard deviation,  $q_e(\text{exp})$  (mg/g) =  $q_e$  obtained from adsorption experiment,  $q_e(\text{cal})$  (mg/g) =  $q_e$  calculated from the pseudo graphs,  $K_1$  ( $\text{min}^{-1}$ ) = rate constant for Pseudo –first order reaction,  $h_0$  is the initial sorption rate (mg/ g min) and  $R^2$ = correlation coefficients, Initial metal ion concentration =100mg/L.

**Figure 7.** Intra-particle diffusion plot for the adsorption process of Cd (II) (geopolymer dosage = 2g/L, Time 20-100 minutes, [Cd (II)] = 100mg/L, T = 298 K, pH 5.0, V = 0.050 L and error bars indicate the average deviation from the mean)

## 4. Conclusions

This study has shown that Kenyan common clays that are of very few economic endeavors may be good raw materials for synthesis of geopolymers for adsorption application of cadmium ions from contaminated water. The adsorbents showed increased the ability to remove Cd (II) ions from synthetic wastewater as the  $\text{SiO}_2/\text{Al}_2\text{O}_3$  ratio increased. The correlation coefficients for the pseudo-second order kinetic model obtained were greater than those of pseudo- first order and therefore indicating that adsorption system studied belonged to the pseudo-second order kinetic model. Thermodynamic studies have shown that the cadmium adsorption process was an endothermic process hence favoured by increase in operation temperatures.

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