

# Comparative Study of Indion-860 and Indion GS-400 Anion Exchange Resins by Application of $^{131}\text{I}$ and $^{82}\text{Br}$ as a Tracer Isotopes

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**Abstract** The present investigation deals with the application of radioactive tracer isotopes  $^{131}\text{I}$  and  $^{82}\text{Br}$  for the performance assessment of two widely used industrial grade ion exchange resins Indion-860 and Indion GS-400. The study reveals that for both the resins, iodide ion-isotopic exchange reaction take place at the faster rate as compared to bromide exchange reaction. Also for both the resins the amount of ions exchanged increases with increase in ionic concentration and decreases with rise in temperature. The above resins were characterized based on kinetics of ion-isotopic exchange reactions. It was observed that the values of specific reaction rate, initial rate of exchange and distribution coefficient values calculated for Indion GS-400 was higher than that calculated for Indion-860 under identical experimental conditions, indicating the superior performance of Indion GS-400 over Indion-860.

**Keywords** Tracer Application, Radioactive Isotopes, Ion-isotopic Exchange Reactions, Reaction Kinetics, Specific Reaction Rate, Distribution Coefficient, Indion GS-400, Indion-860

## 1. Introduction

Radioactive tracers were first applied to industrial problem solving around the middle of the last century. Since then their use has increased steadily so that, at present, radiotracer techniques are used extensively throughout the world for troubleshooting and process optimization in industry. The economic benefits that may be derived from the use of this technology are great, a fact that is recognized by the governments of developing countries. Though radioisotopes have been applied to the solution of problems in industry for over 50 years, research and development of the technology continues unabated. There are two main reasons for the continuing interest. Firstly, it is industry driven. Because of their unique properties, radioactive isotopes can be used to obtain information about plants and processes that cannot be obtained in any other way. Often, the information is obtained with the plant on-stream and without disrupting the process in any way. This can lead to substantial economic benefits, from shutdown avoidance to process optimization. Secondly, the methodology is derived from many fields of science and technology including radioisotope production, radiation detection, data acquisition, treatment and analysis, and mathematical modeling. The

radioisotopes have proved as a tool to study many problems in chemical, biological and medicinal fields. Radiotracers have helped in identification of leaks in buried pipelines and dams. Process parameters such as mixing efficiency, residence time, flow rate, material inventory and silt movement in harbours are studied using radioisotopes[1]. The efficiency of several devices in a wastewater treatment plant (primary and secondary clarifiers, aeration tank) is also investigated by means of radiotracers[2].

The fundamental principle in radiochemical investigations is that the chemical properties of a radioisotope of an element are almost the same as those of the other stable/radioactive isotopes of the element. When radioisotope is present in a chemical form identical to that of the bulk of the element in a chemical process, then any reaction the element undergoes can be directly traced by monitoring the radioisotope. Radiochemical work involves two main steps first is the sampling of chemical species to be studied and second is quantitative determination of the radiation emitted by the radioisotope in the sample[1]. In radiotracer study, a short lived radioisotope in a physico-chemical form similar to that of the process material is used to trace the material under study. The radioisotopes in suitable physical and chemical forms are introduced in systems under study. By monitoring the radioactivity both continuously or after sampling (depending on the nature of study), the movement, adsorption, retention etc. of the tracer and in turn, of the bulk matter under investigation, can be followed. The tracer concentration recorded at various locations also helps to

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draw information about the dynamic behavior of the system under study. The radioisotopes preferred for such studies are gamma emitters having half-life compatible with the duration of studies. The strength of radioactivity used varies depending on the nature of application. Applications of radiotracers in chemical research cover the studies of reaction mechanism, kinetics, exchange processes and analytical applications such as radiometric titrations, solubility product estimation, isotope dilution analysis and autoradiography.

Considering the above wide use of radioactive isotopes in various industrial and technical applications, in the present investigation, they are applied to assess the performance of industrial grade anion exchange resins Indion-860 and Indion GS-400 under different operational parameters like temperature and ionic concentrations. It is expected that the tracer technique used here can also be used for characterization of other organic ion exchange resins which are synthesized for their specific technical applications[3-7]. The present technique can also be extended further to standardize the operational parameters so as to bring about the most efficient performance of those resins in their specific industrial applications.

## 2. Experimental

### 2.1. Conditioning of Ion Exchange Resins

**Table 1.** Properties of ion exchange resins

Ion exchange resin	Matrix	Particle Size (mm)	Moisture content (%)	Operating pH	Maximum operating Temperature ( $^{\circ}\text{C}$ )	Total exchange capacity (meq./mL)
Indion-860	Crosslinked Polystyrene	0.3-1.2	52	0-7	60	1.4
Indion GS-400	Crosslinked Polystyrene	0.3-1.2	54	0-14	75	1.6

Ion exchange resin Indion-860 is a weakly basic macroporous resin having  $-\text{N}^+\text{R}_2$  functional group; while Indion GS-400 is a gel type strongly basic anion exchange resin having  $-\text{N}^+\text{R}_3$  functional group. Both the resins as supplied by the manufacturer (Ion exchange India Ltd., Mumbai) were in chloride form. Details regarding the properties of the resins used are given in Table 1. These resins were converted separately in to iodide / bromide form by treatment with 10 % KI / KBr solution in a conditioning column which is adjusted at the flow rate as 1 mL / min. The resins were then washed with double distilled water, until the

washings were free from iodide/bromide ions as tested by  $\text{AgNO}_3$  solution. These resins in bromide and iodide form were then dried separately over  $\text{P}_2\text{O}_5$  in desiccators at room temperature.

### 2.2. Radioactive Tracer Isotopes

The radioisotope  $^{131}\text{I}$  and  $^{82}\text{Br}$  used in the present experimental work was obtained from Board of Radiation and Isotope Technology (BRIT), Mumbai. Details regarding the isotopes used in the present experimental work are given in Table 2.

**Table 2.** Properties of  $^{131}\text{I}$  and  $^{82}\text{Br}$  tracer isotopes[8]

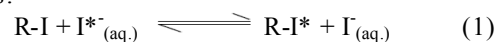
Isotopes	Half-life	Radioactivity / mCi	$\gamma$ -energy / MeV	Chemical form	Physical form
$^{131}\text{I}$	8.04 d	5	0.36	Iodide*	Aqueous
$^{82}\text{Br}$	36 h	5	0.55	Bromide**	Aqueous

\* Sodium iodide in dilute sodium sulphite

\*\* Ammonium bromide in dilute ammonium hydroxide

### 2.3. Study on Kinetics of Iodide Ion-isotopic Exchange Reaction

In a stoppered bottle 250 mL ( $V$ ) of 0.001 M iodide ion solution was labeled with diluted  $^{131}\text{I}$  radioactive solution using a micro syringe, such that 1.0 mL of labeled solution has a radioactivity of around 15,000 cpm (counts per minute) when measured with  $\gamma$ -ray spectrometer having NaI (TI) scintillation detector. Since only about 50–100  $\mu\text{L}$  of the radioactive iodide ion solution was required for labeling the solution, its concentration will remain unchanged, which was further confirmed by potentiometer titration against  $\text{AgNO}_3$  solution. The above labeled solution of known initial activity ( $A_i$ ) was kept in a thermostat adjusted to  $30.0^{\circ}\text{C}$ . The swelled and conditioned dry ion exchange resins in iodide form weighing exactly 1.000 g ( $m$ ) were transferred quickly into this labeled solution which was vigorously stirred by using mechanical stirrer and the activity in cpm of 1.0 mL of solution was measured. The solution was transferred back to the same bottle containing labeled solution after measuring activity. The iodide ion-isotopic exchange reaction can be represented as:



Here R-I represents ion exchange resin in iodide form;  $\text{I}^*_{(\text{aq.})}$  represents aqueous iodide ion solution labeled with  $^{131}\text{I}$  radiotracer isotope.

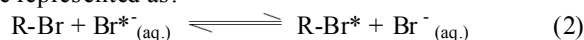
The activity of solution was measured at a fixed interval of every 2.0 min. The final activity ( $A_f$ ) of the solution was also measured after 3h which was sufficient time to attain the equilibrium[9-14]. The activity measured at various time intervals was corrected for background counts.

Similar experiments were carried out by equilibrating

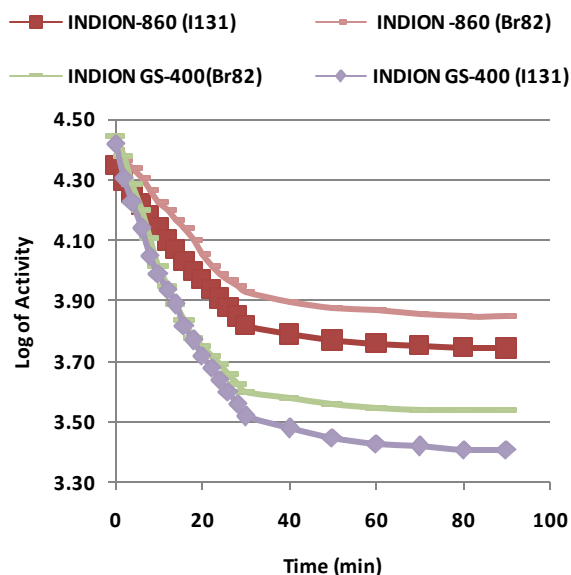
separately 1.000 g of ion exchange resin in iodide form with labeled iodide ion solution of four different concentrations ranging up to 0.004 M at a constant temperature of 30.0°C. The same experimental sets were repeated for higher temperatures up to 45.0°C.

#### 2.4. Study on Kinetics of Bromide Ion-isotopic Exchange Reaction

The experiment was also performed to study the kinetics of bromide ion- isotopic exchange reaction by equilibrating 1.000 g of ion exchange resin in bromide form with labeled bromide ion solution in the same concentration and temperature range as above. The labeling of bromide ion solution was done by using  $^{82}\text{Br}$  as a radioactive tracer isotope for which the same procedure as explained above was followed. The bromide ion-isotopic exchange reaction can be represented as:

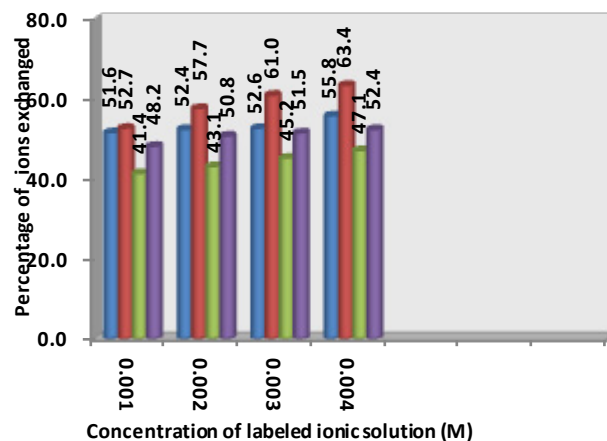


Here R-Br represents ion exchange resin in bromide form;  $\text{Br}^{*-}_{(\text{aq.})}$  represents aqueous bromide ion solution labeled with  $^{82}\text{Br}$  radiotracer isotope.



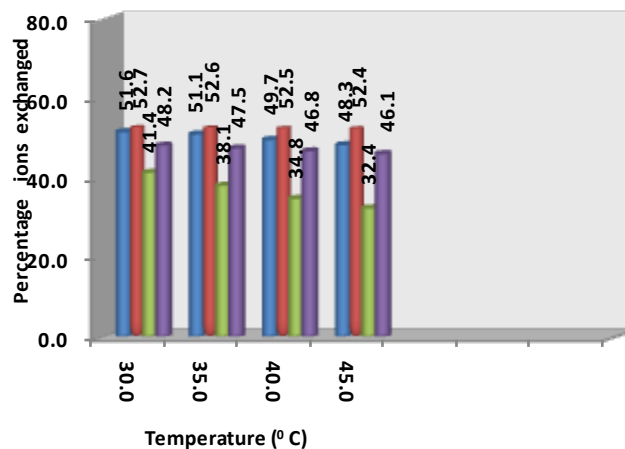
**Figure 1.** Kinetics of Ion-Isotopic Exchange Reactions Amount of ion exchange resin = 1.000 g, Concentration of labeled exchangeable ionic solution = 0.001M, Volume of labeled ionic solution = 250 mL, Temperature = 30.0°C

■ INDION-860 (Reaction 1) ■ INDION GS-400 (Reaction 1)  
■ INDION-860 (Reaction 2) ■ INDION GS-400 (Reaction 2)

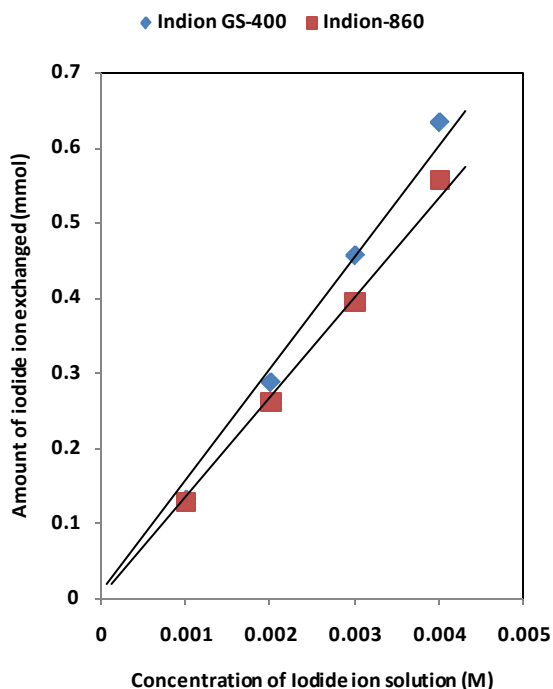


**Figure 2.** Variation in Percentage Ions Exchanged with Concentration of Labeled Ionic Solution Amount of ion exchange resin = 1.000 g, Volume of labeled ionic solution = 250 mL, Temperature = 30.0°C

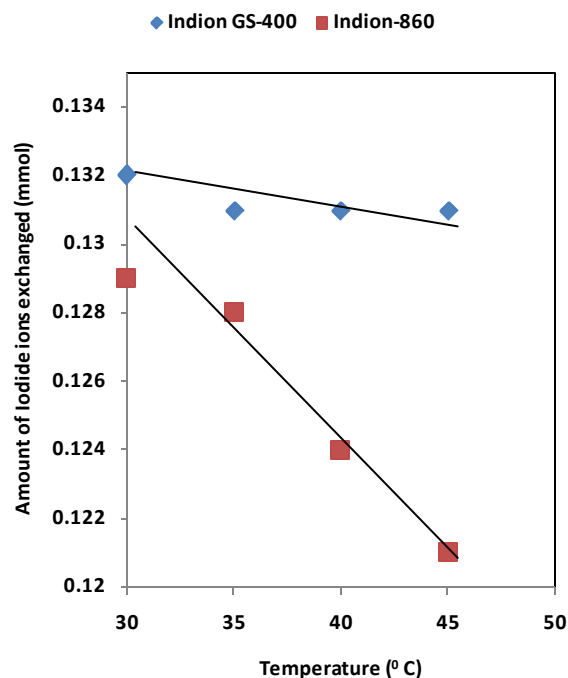
■ INDION-860 (Reaction 1) ■ INDION GS-400 (Reaction 1)  
■ INDION-860 (Reaction 2) ■ INDION GS-400 (Reaction 2)



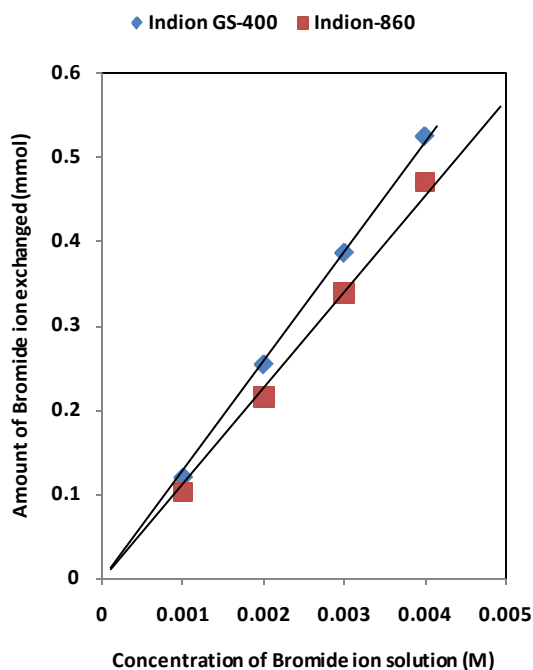
**Figure 3.** Variation in Percentage Ions Exchanged with Temperature of Labeled Ionic Solution Amount of ion exchange resin = 1.000 g, Concentration of labeled exchangeable ionic solution = 0.001M, Volume of labeled ionic solution = 250 mL, Amount of exchangeable ions in 250 mL labeled solution = 0.250 mmol



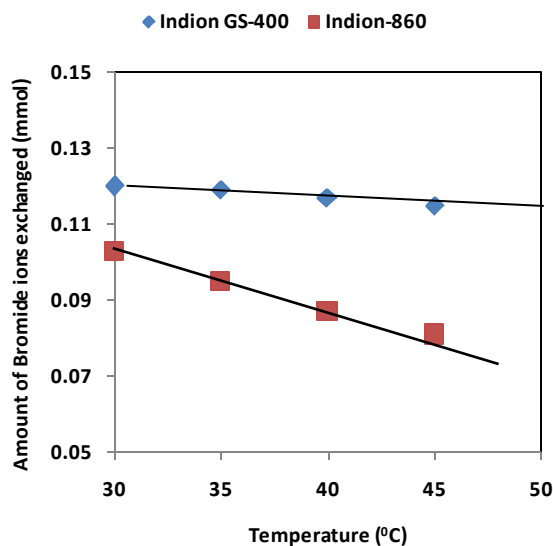
**Figure 4.** Correlation between concentrations of iodide ion solution and amount of iodide ion exchanged Amount of ion exchange resin = 1.000 g, Volume of labeled ionic solution = 250 mL, Temperature = 30.0°C Correlation coefficient (r) for Indion-860 = 0.9987 Correlation coefficient (r) for Indion GS-400 = 0.9996



**Figure 6.** Correlation between Temperatures of exchanging medium and amount of iodide ion exchanged Amount of ion exchange resin = 1.000 g, Concentration of labeled exchangeable ionic solution = 0.001M, Volume of labeled ionic solution = 250 mL, Amount of exchangeable ions in 250 mL labeled solution = 0.250 mmol Correlation coefficient (r) for Indion-860 = -0.9778 Correlation coefficient (r) for Indion GS-400 = -0.7746



**Figure 5.** Correlation between concentrations of bromide ion solution and amount of bromide ion exchanged Amount of ion exchange resin = 1.000 g, Volume of labeled ionic solution = 250 mL, Temperature = 30.0°C Correlation coefficient (r) for Indion-860 = 0.9994 Correlation coefficient (r) for Indion GS-400 = 1.0000



**Figure 7.** Correlation between Temperatures of exchanging medium and amount of bromide ion exchanged Amount of ion exchange resin = 1.000 g, Concentration of labeled exchangeable ionic solution = 0.001M, Volume of labeled ionic solution = 250 mL, Amount of exchangeable ions in 250 mL labeled solution = 0.250 mmol Correlation coefficient (r) for Indion-860 = -0.9978 Correlation coefficient (r) for Indion GS-400 = -0.9898

**Table 3.** Concentration effect on Ion-Isotopic Exchange Reactions Amount of ion exchange resin = 1.000 g, Volume of labeled ionic solution = 250 mL, Temperature = 30.0°C

Concentration of ionic solution (M)	Amount of ions in 200 mL solution (mmol)	REACTION -1								REACTION -2							
		INDION-860				INDION GS-400				INDION-860				INDION GS-400			
		Specific reaction rate of rapid process $\text{min}^{-1}$	Amount of iodide ion exchanged (mmol)	Initial rate of iodide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid process $\text{min}^{-1}$	Amount of iodide ion exchanged (mmol)	Initial rate of iodide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid process $\text{min}^{-1}$	Amount of bromide ion exchanged (mmol)	Initial rate of bromide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid process $\text{min}^{-1}$	Amount of bromide ion exchanged (mmol)	Initial rate of bromide ion exchange (mmol/min)	Log $K_d$
0.001	0.250	0.183	0.129	0.024	17.0	0.199	0.132	0.026	17.4	0.130	0.103	0.013	12.0	0.153	0.120	0.018	12.5
0.002	0.500	0.190	0.262	0.050	17.9	0.215	0.288	0.062	18.4	0.136	0.216	0.029	12.3	0.169	0.254	0.043	12.9
0.003	0.750	0.198	0.395	0.078	18.6	0.230	0.457	0.105	19.2	0.143	0.339	0.048	13.1	0.181	0.386	0.070	13.7
0.004	1.000	0.209	0.558	0.117	19.3	0.245	0.634	0.155	19.9	0.149	0.471	0.070	13.9	0.190	0.524	0.100	14.4

**Table 4.** Temperature effect on Ion-Isotopic Exchange Reactions Amount of ion exchange resin = 1.000 g, Concentration of labeled exchangeable ionic solution = 0.001M, Volume of labeled ionic solution = 250 mL, Amount of exchangeable ions in 250 mL labeled solution = 0.25 mmol

Temperature °C	REACTION -1								REACTION -2							
	INDION-860				INDION GS-400				INDION-860				INDION GS-400			
	Specific reaction rate of rapid Process $\text{min}^{-1}$	Amount of iodide ion exchanged (mmol)	Initial rate of iodide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid Process $\text{min}^{-1}$	Amount of iodide ion Exchanged (mmol)	Initial rate of iodide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid Process $\text{min}^{-1}$	Amount of bromide ion Exchanged (mmol)	Initial rate of bromide ion exchange (mmol/min)	Log $K_d$	Specific reaction rate of rapid Process $\text{min}^{-1}$	Amount of bromide ion exchanged (mmol)	Initial rate of bromide ion exchange (mmol/min)	Log $K_d$
30.0	0.183	0.129	0.024	17.0	0.199	0.132	0.026	17.4	0.130	0.103	0.013	12.0	0.153	0.120	0.018	12.5
35.0	0.171	0.128	0.022	16.4	0.196	0.131	0.026	16.9	0.119	0.095	0.011	11.6	0.150	0.119	0.018	12.0
40.0	0.162	0.124	0.020	15.9	0.192	0.131	0.025	16.4	0.108	0.087	0.009	10.9	0.148	0.117	0.017	11.4
45.0	0.154	0.121	0.019	15.4	0.190	0.131	0.025	16.0	0.100	0.081	0.008	10.2	0.146	0.115	0.017	10.7

### 3. Results and Discussion

#### 3.1. Comparative Study of Ion-isotopic Exchange Reactions

In the present investigation it was observed that due to the rapid ion-isotopic exchange reaction taking place, the activity of solution decreases rapidly initially, then due to the slow exchange the activity of the solution decreases slowly and finally remains nearly constant. Preliminary studies show that the above exchange reactions are of first order [15,

16]. Therefore logarithm of activity when plotted against time gives a composite curve in which the activity initially decreases sharply and thereafter very slowly giving nearly straight line (Figure 1), evidently rapid and slow ion-isotopic exchange reactions were occurring simultaneously [9-14]. Now the straight line was extrapolated back to zero time. The extrapolated portion represents the contribution of slow process to the total activity which now includes rapid process also. The activity due to slow process was subtracted from the total activity at various time intervals. The difference

gives the activity due to rapid process only. From the activity exchanged due to rapid process at various time intervals, the specific reaction rates ( $k$ ) of rapid ion-isotopic exchange reaction were calculated. The amount of iodide / bromide ions exchanged (mmol) on the resin were obtained from the initial and final activity of solution and the amount of exchangeable ions in 250 mL of solution. From the amount of ions exchanged on the resin (mmol) and the specific reaction rates ( $\text{min}^{-1}$ ), the initial rate of ion exchanged ( $\text{mmol/min}$ ) was calculated.

Because of larger solvated size of bromide ions (310 pm) as compared to that of iodide ions (300 pm), it was observed that the exchange of bromide ions occurs at the slower rate than that of iodide ions [17]. Hence under identical experimental conditions, the values of specific reaction rate ( $\text{min}^{-1}$ ), amount of ion exchanged (mmol) and initial rate of ion exchange ( $\text{mmol/min}$ ) are calculated to be lower for bromide ion-isotopic exchange reaction than that for iodide ion-isotopic exchange reaction as summarized in Tables 3 and 4. For both bromide and iodide ion-isotopic exchange reactions, under identical experimental conditions, the values of specific reaction rate increases with increase in concentration of ionic solution from 0.001M to 0.004M (Table 3). However, with rise in temperature from  $30.0^\circ\text{C}$  to  $45.0^\circ\text{C}$ , the specific reaction rate was observed to decrease (Table 4). From the results, it appears that iodide ions exchange at the faster rate as compared to that of bromide ions which was related to the extent of solvation (Tables 3 and 4).

From the knowledge of  $A_i$ ,  $A_f$ , volume of the exchangeable ionic solution ( $V$ ) and mass of ion exchange resin ( $m$ ), the  $K_d$  value was calculated by the equation

$$K_d = [(A_i - A_f) / A_f] \times V / m \quad (3)$$

Heumann *et al.* [18] in the study of chloride distribution coefficient on strongly basic anion exchange resin observed that the selectivity coefficient between halide ions increased at higher electrolyte concentrations. Adachi *et al.* [19] observed that the swelling pressure of the resin decreased at higher solute concentrations resulting in larger  $K_d$  values. The temperature dependence of  $K_d$  values on cation exchange resin was studied by Shuji *et al.* [20]; were they observed that the values of  $K_d$  increased with fall in temperature. The present experimental results also indicates that the  $K_d$  values for bromide and iodide ions increases with increase in ionic concentration of the external solution, however with rise in temperature the  $K_d$  values were found to decrease. It was also observed that the  $K_d$  values for iodide ion-isotopic reaction were calculated to be higher than that for bromide ion-isotopic reaction (Tables 3 and 4).

### 3.2. Comparative Study of Anion Exchange Resins

From the Table 3, it is observed that for iodide ion-isotopic exchange reaction by using Indion GS-400 resin, the values of specific reaction rate ( $\text{min}^{-1}$ ), amount of iodide ion exchanged (mmol), initial rate of iodide ion exchange ( $\text{mmol/min}$ ) and  $\log K_d$  were 0.199, 0.132, 0.026 and 17.4

respectively, which was higher than 0.183, 0.129, 0.024 and 17.0 respectively as that obtained by using Indion-860 resins under identical experimental conditions of  $30.0^\circ\text{C}$ , 1.000 g of ion exchange resins and 0.001 M labeled iodide ion solution. The identical trend was observed for the two resins during bromide ion-isotopic exchange reaction.

From Table 3, it is observed that using Indion GS-400 resins, at a constant temperature of  $30.0^\circ\text{C}$ , as the concentration of labeled iodide ion solution increases 0.001 M to 0.004 M, the percentage of iodide ions exchanged increases from 52.7 % to 63.4 %. While using Indion-860 resins under identical experimental conditions the percentage of iodide ions exchanged increases from 51.6 % to 55.8 %. Similarly in case of bromide ion-isotopic exchange reaction, the percentage of bromide ions exchanged increases from 48.2 % to 52.4 % using Indion GS-400 resin, while for Indion-860 resin it increases from 41.4 % to 47.1 %. The effect of ionic concentration on percentage of ions exchanged is graphically represented in Figure 2.

From Table 4, it is observed that using Indion GS-400 resins, for 0.001 M labeled iodide ion solution, as the temperature increases  $30.0^\circ\text{C}$  to  $45.0^\circ\text{C}$ , the percentage of iodide ions exchanged decreases from 52.7 % to 52.4 %. While using Indion-860 resins under identical experimental conditions the percentage of iodide ions exchanged decreases from 51.6 % to 48.3 %. Similarly in case of bromide ion-isotopic exchange reaction, the percentage of bromide ions exchanged decreases from 48.2 % to 46.1% using Indion GS-400 resin, while for Indion-860 resin it decreases from 41.4 % to 32.4 %. The effect of temperature on percentage of ions exchanged is graphically represented in Figure 3.

The overall results indicate that under identical experimental conditions, as compared to Indion-860 resins, Indion GS-400 resins shows higher percentage of ions exchanged. Thus Indion GS-400 resins show superior performance than Indion-860 resins.

### 3.3. Statistical Correlations

The results of present investigation show a strong positive linear co-relationship between amount of ions exchanged and concentration of ionic solution (Figures 4, 5). In case of iodide ion-isotopic exchange using Indion GS-400 and Indion-860 resins, the values of correlation coefficient ( $r$ ) were found to be 0.9996 and 0.9987 respectively, while for bromide ion-isotopic exchange the respective values of  $r$  were calculated as 1.0000 and 0.9994.

There also exist a negative co-relationship between amount of ions exchanged and temperature of exchanging medium (Figures 6, 7). For Indion GS-400 during iodide isotopic exchange reaction somewhat poor correlation was observed having value of ( $r$ ) -0.7746, however for Indion-860 resins, excellent correlation exist having ( $r$ ) value of -0.9778. Similarly in case of bromide ion-isotopic exchange the ( $r$ ) values calculated for Indion GS-400 and

Indion-860 were -0.9898 and -0.9978 respectively.

## 4. Conclusions

The experimental work carried out in the present investigation will help to standardize the operational process parameters so as to improve the performance of selected ion exchange resins. The radioactive tracer technique used here can also be applied for characterization of different nuclear as well as non-nuclear grade ion exchange resins.

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