

Nano Fibrous Cerium (IV) Hydrogen Phosphate Membrane Self Supported Aniline Polymerization Agent

S. K. Shakshooki*, M. A. Abuain, Amani M. El-Aouzi

Department of Chemistry, Faculty of Science, Tripoli University, Tripoli, Libya

Abstract Nanosized fibrous cerium (iv) Hydrogen phosphate membrane, $\text{Ce}(\text{HPO}_4)_2 \cdot 9\text{H}_2\text{O}$ (nCeP_f), was prepared and characterized. Novel supported nanofibrous cerium phosphate / polyaniline nanocomposite membranes were prepared via in-situ chemical oxidation of the monomer aniline, that was promoted by the reduction of $\text{Ce}(\text{iv})$ ions present in the inorganic matrix. The presence of $\text{Ce}(\text{iv})$ ions allows redox reactions necessary to oxidative polymerization to occur. The resultant materials were characterized by TGA, elemental (C, H, N, S) analysis and FT-IR spectroscopy. Images SEM of the resulting nanocomposites reveal a uniform distribution of the polymer on the inorganic matrix. A possible explanation is nCeP_f present on the surface of the fiber is attacked by the aniline and converted to cerium(III) orthophosphate (CePO_4). The rest of the (nCeP_f) was preserved. The degree of polymer loading and its type can be controlled by varying method of the preparation. Accordingly two types of green emeraldine salts (in form of chloride and sulfate salts) were obtained. Amount of polyaniline present in the resultant composites found to be in the range of (4.5 - 16.74% in weight).

Keywords Nanofibrous cerium (iv) hydrogen phosphate membrane, Aniline polymerization

1. Introduction

Conducting polymers are a novel class of synthetic metals that combine the chemical, electrochemical and mechanical properties of polymers with the electronic properties of metals and semiconductor, generated tremendous interest due to their potential applications in various fields such as rechargeable batteries, electrochromic display devices, separation membranes, sensors and anticorrosive coatings on metals [1-11]. Heterocyclic conducting polymers containing nitrogen atoms like polyaniline, polyindole and polypyrrole, and their substituted derivatives have received increasing attention in various fields, electronics, industry and others. Their electrical and electrochemical properties show great promise for commercial applications. They provide an interesting and useful focus for expansion of the fields of polymeric reagent research and molecular engineering. They are electrically conductive which makes their use as metal replacement materials of some interest [12-16].

The properties of electronically conductive polymers depend strongly on their microstructure and morphology, and these determined by the synthesis method, counter ion and several other variables [17].

Polyaniline is one of the most studied conducting polymers, has potential for applications due to its light

weight, conductivity, mechanical flexibility and low cost. Polyaniline is especially attractive because it is relatively inexpensive. It has certain advantages over other conducting polymers, including simplicity and rapidity of preparation by electrochemical and oxidative chemical polymerization methods, and the ability to be formed in aqueous electrolytic solutions. It has good environmental stability and its conductivity could be controlled with acid/base (doping/undoping) [12, 13, 17-30].

Inorganic layered nanomaterials are receiving great attention because of their size, structure, and possible biochemical applications [31, 32], that have been proven to be good carriers for organic polar molecules. Examples of these are zirconium phosphates. Taking advantage of the expandable interlayer space of the layered materials. Researchers have been capable of encapsulating functional biomolecules into these inorganic matrices protecting them from interacting with environment, avoiding denaturation and enhancing their shelf life [31, 32].

Crystalline cerium phosphates have been studied for a long time as ion exchangers, their structures remains unknown until recently [33-35]. The reason is that, the composition, the structure and the degree of crystallinity of their precipitates results from reaction of solutions containing a $\text{Ce}(\text{iv})$ salt is mixed with a solution of phosphoric acid of $[(\text{PO}_4) / \text{CeIV ratio}]$, strongly dependent on the experimental conditions such as rate and order of mixing of the solutions, stirring, temperature and digestion time, this also implemented on fibrous cerium phosphate [36]. To date most of the work on fibrous cerium phosphate

* Corresponding author:

shakshooki2002@yahoo.com (S. K. Shakshooki)

Published online at <http://journal.sapub.org/chemistry>

Copyright © 2015 Scientific & Academic Publishing. All Rights Reserved

was carried out on its ion exchange [37], intercalation [38] and electrical conductance properties [39]. Studies on its polystyrene, polyacrylamide [40] and (polyvinyl chloride-based polyvinyl (alcohol) [41] composites have been reported. Nanoscaled tetravalent metal phosphates and their organic polymer composites comprise an important class of synthetic engineering. However; research in such area is still its infancy [42-45]. Nanotechnologies are at the center of numerous investigations and huge investments. However, chemistry has anticipated for long the importance decreasing the size in the search of new properties of materials, and of materials structured at the nanosize in a number of applications relate to daily life. Organic-inorganic nanocomposite membranes have gained great attention recently [45, 46], The composite material may combine the advantage of each material, for instance, flexibility, processability of polymers and the selectivity and thermal stability of the inorganic filler [43-47]. Pyrrole polymerization by fibrous cerium phosphate has been reported [48, 49]. polymerization of Pyrrole and aniline on α -Tin phosphate and copper zirconium phosphate was reported [50, 51].

In our laboratory we are carrying systematic investigations on novel tetravalent metal phosphates / organic heterocyclic conducting polymers nanocomposite membranes. Recently we have reported [52, 53] the preparation and characterization of fibrous cerium phosphate/polybenzimidazole nanocomposite membrane and fibrous cerium phosphate / polyindole nanocomposite membranes. The present study describes the preparation and characterization of novel supported fibrous cerium phosphate / polyaniline nanocomposite membranes via in-situ chemical oxidation, of the monomer aniline, that was promoted by the reduction of Ce(IV) ions present in the inorganic matrix.

2. Materials and Methods

2.1. Chemicals

Ce(SO₄)₂·4H₂O, H₃PO₄(85%) of BDH, aniline (99.5%) of Mindex UK. Other reagents used were of analytical grade.

2.2. Instruments Used for Characterization

X-Ray powder diffractometer Siemens D-500, using Ni-filtered CuK α (λ = 1.54056Å), TG/DTA SHIExtra6000 TGA Perkin Elmer thermogravimetric analyzer(TGA7)US, CHNS-Elemental Analysis, Vario Elementa German. Fourier Transform IR spectrometer, model FT/IR-6100, Scanning electron microscopy (SEM) Jeol SMJ Sm 5610 LV.

2.3. Preparation of Nanofibrous Cerium Phosphate Membrane, Ce(HPO₄)₂·2.9H₂O

Nanofibrous cerium phosphate membrane was prepared from adding 300 ml of 0.05M CeSO₄·4H₂O in 0.5 M H₂SO₄

solution, drop wise, to 300 ml of 6 M H₃PO₄ at ~ 80°C with stirring. After complete addition the resultant material left to digest at that temperature for 4h. To that 3 liter of hot distilled water, (~60°C), was added with stirring for 1h. The resulting fibrous cerium phosphate was subjected to washing with distilled water up to pH 3, then filtered and left to dry in air.

2.4. Exchange Capacity

Exchange capacity of the resultant nanosized fibrous cerium phosphate membrane was determined by addition of 25 ml of 0.1M NaCl solution to 100 mg of the material, with stirring for one h, then titrated with 0.1 M NaOH solution.

2.5. Thermal Analysis

Thermal analyses were carried out at temperature range about 20 - 775°C in nitrogen atmosphere, the rate was 10°C/min.

2.6. Preparation of Nanofibrous Cerium Phosphate Polyaniline Composite Membranes

Cerium phosphate / polyaniline nanocomposite membranes were prepared by different methods:

- (1) By immersion of fibrous cerium phosphate self supported-sheet 160 mg in 10 ml ethanol solution containing 180 mg aniline in 10 ml ethanol solution, at room temperature, for 48h. The impregnated sheet was removed, washed with ethanol and left to dry in air. The color of the resulting product was dark-blackish-blue, designated as composite(I).
- (2) By immersion of fibrous cerium phosphate self supported-sheet 160 mg in 10 ml ethanol solution containing 180 mg aniline at room temperature for 1h. The impregnated sheet was removed, washed with ethanol and left to dry in air. The color of the resulting product was pale-blue, designated as composite(II).
- (3) immersion of fibrous cerium phosphate self supported-sheet 160 mg in 10 ml 1M HCl solution for 24h, resulting impregnated sheet of pale green in color, it was washed with ethanol and left to dry in air, and designated as composite(III).
- (4) Addition of 0.5ml aniline to 35 ml of the original slurry aqueous solution of fibrous cerium phosphate (prior subjecting to washing, its pH =1.6), with stirring at room temperature leading to immediate formation of green polyaniline, (emeraldine salt). The stirring of the mixture was continued for an hour. The contents left in fridge (5°C) for 24h. The product was filtered washed with ethanol and left to dry in air. The color of the resultant product dark green, designated as composite (IV).

Nanofibrous cerium phosphate membrane, Ce(HPO₄)₂·2.9H₂O (nCePf), was prepared and characterized by chemical, XRD, TGA, FT-IR, and SEM.

1. XRD

XRD of (nCePf) is shown in Figure 1, with $d_{001} = 10.89\text{\AA}$.

2. TGA

Thermogram of (nCePf) is shown in Figure 2. The thermal decomposition occurs in continuous process, the thermal analysis was carried out at temperatures between 10-775°C, the final product was CeP_2O_7 . Loss of water of hydration occurs between 60-200°C, followed by POH groups condensation. The total weight loss found to be equal to 19.09%.

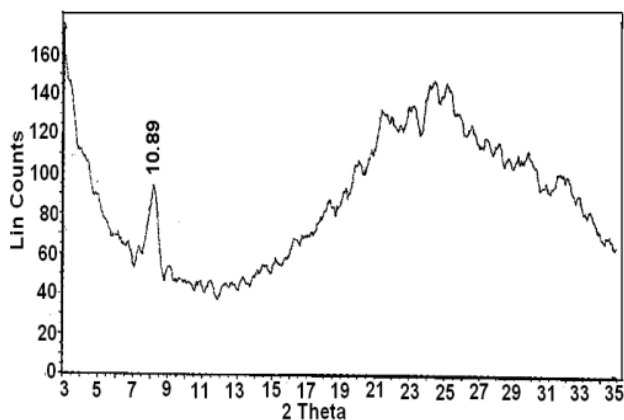


Figure 1. XRD of (nCePf)

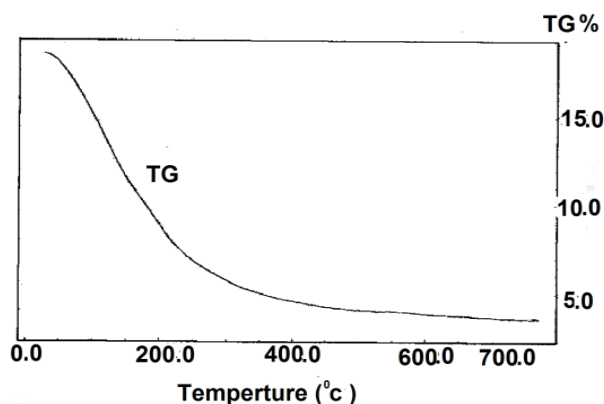


Figure 2. TGA of (nCePf)

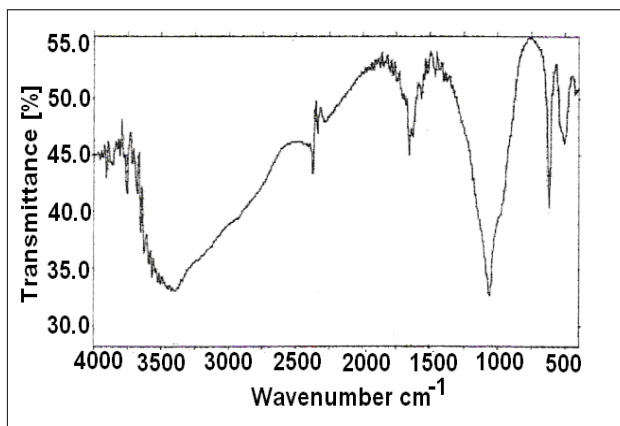


Figure 3. FT-IR spectra of (nCePf)

3. FT-IR

Fig. 3 shows FT-IR spectrum of fibrous $\text{Ce}(\text{HPO}_4)_2 \cdot 2.9\text{H}_2\text{O}$, with a trend similar to that of M(IV) phosphates. It consists of broad band centered at 3350cm^{-1} is due to OH groups symmetric stretching of H_2O , small sharp band at 1628cm^{-1} is related to H-O-H bending, and sharp broad band centered at 1045cm^{-1} is corresponds to phosphate groups vibration. The bands at the region $630\text{--}450\text{ cm}^{-1}$ are ascribe the presence of $\delta(\text{PO}_4)$.

4. SEM

SEM morphology image of the nanosized fibrous cerium phosphate (nCePf) is shown in Figure 4. The photograph shows its average size is $\sim 20.5\text{ nm}$.

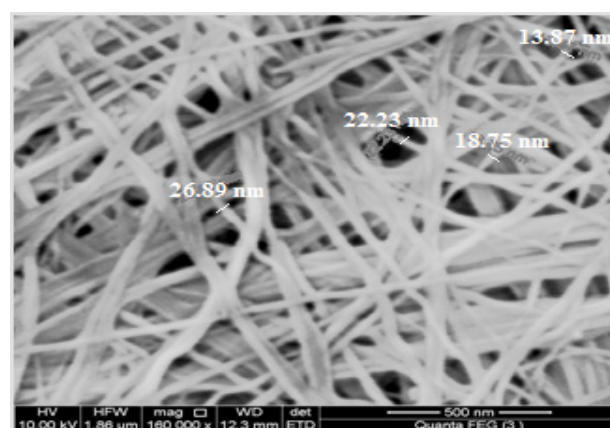


Figure 4. SEM morphology image of (nCePf)

5. Ion Exchange Capacity

The ion exchange capacity of nanosized fibrous cerium(IV) phosphate, $\text{Ce}(\text{HPO}_4)_2 \cdot 2.9\text{H}_2\text{O}$, found to be equal to 5.21 meq/g.

New supported polyaniline nanocomposites were prepared and investigated in this work. Self-supported sheet of nanosized fibrous cerium (IV) hydrogenphosphate, $\text{Ce}(\text{HPO}_4)_2 \cdot 2.9\text{H}_2\text{O}$ (nCePf), was used as the inorganic matrix, since the fibrous morphology present by this material makes molding possible, giving desired shape to the final (nCePf)/polyaniline nanocomposite membranes.

It was found that when nanosized fibrous cerium(IV) hydrogen phosphate sheet was immersed in ethanolic solution of aniline the color of self supported sheet gradually changes with time to (pale blue and finally to dark blackish-blue) due to the formation of nanocomposite membrane (composite I).

When nanosized fibrous cerium(IV) hydrogen phosphate sheet was immersed in ethanolic solution of aniline the color of self supported sheet changes immediately to pale blue (composite II).

On immersing of nanosized fibrous cerium(IV) hydrogen phosphate sheet in aniline HCl solution the color of self supported sheet gradually changes with time to pale blue and finally to light green (composite III).

However, when ethanolic solution was added to the

original slurry solution of fibrous cerium phosphate (prior subjecting to washing, pH1.6), on filtration resulting in formation of nanocomposite membrane, dark green in color (composite IV). The color changes can be considered the first evidence of polymer formation. In the resulting composites the self-supported sheet integrity is preserved. This result is interesting because the shape integrity makes the building of molded conducting device possible. The resultant composites were characterized by elemental (C, H, N, S) analysis, TGA, FT-IR and SEM.

6. Elemental Analysis

From elemental (C, H, N, S) analysis, the amount of organic material present in the (nCePf)/ PANI composite(I), blackish-blue in color, was 7.759% in weight (where 7.89% in weight from TGA analysis). The % of elemental (C, H, N) related to the organic polymer found to be C = 6.33, H =

0.476% and N = 0.953%. The elemental (C, H, N) analysis of pale blue composite (II), obtained from immersion of fibrous cerium phosphate in aniline ethanolic solution for a period < 20 min, was 4.462% in weight. Its elemental (C, H, N) analysis were C=3.65, H=0.273, and N=0.549% in weight. The elemental (C, H, N) analysis related for the organic polymer of pale green composite (III) were 11.695% in weight, where C=9.244, H=0.706, and N=1.745% in weight. Reaction of aniline with fibrous cerium phosphate in dilute sulfuric acid solution lead to immediate formation of green polyaniline, (emeraldine salt). The elemental (C, H, N, S) analysis related for the organic polymer of dark-green composite (IV) found to be 16.74% in weight, where C=12.69, H=1.0, and N=2.44% in weight, S=2.05% in weight. The degree of polymer loading can be controlled by varying method of the preparation and aging time.

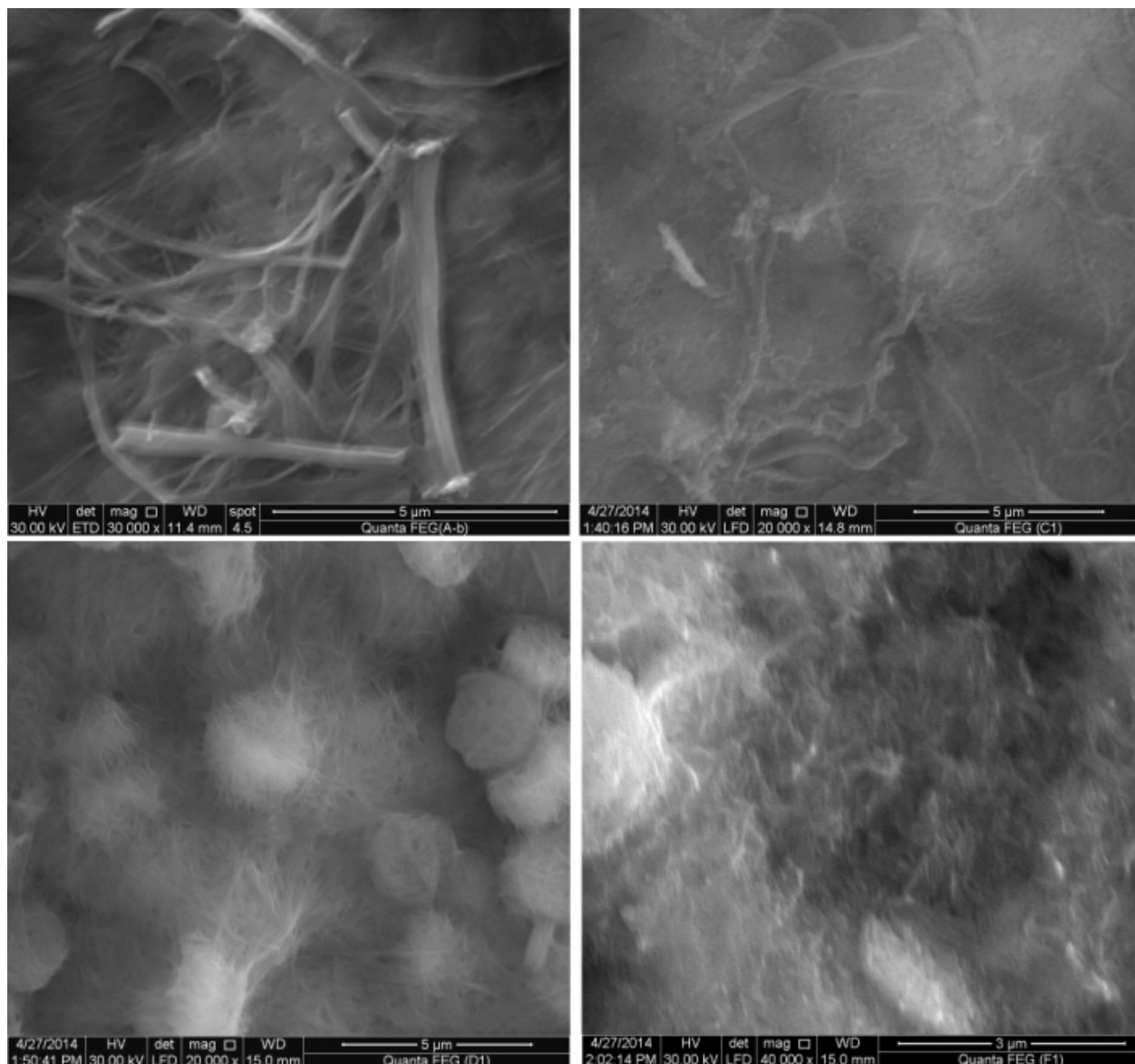


Figure 5. SEM images of (nCePf)/PANI nanocomposites (I-IV) respectively

7. SEM

We have noticed, from SEM analysis, the nanosized fibrous cerium phosphate morphology was preserved, consequently physical integrity of the resulting nanocomposites were maintained. Figure (5) shows scanning electron microscopy (SEM) images for (nCePf) / polyaniline nanocomposites (I), (II), (III) and (VI), respectively, reveal a uniform distribution of the polymer on the inorganic matrix, coating the nCePf and filling the empty spaces between the fibers. In spite of the likely conversion of some of nCePf to CePO_4 , nanocomposites present fibrous morphology. A possible explanation is nCePf present on the surface of the fiber is attacked by the aniline and converted to cerium (III) orthophosphate. The rest of the nCePf was preserved.

8. FT-IR of Nanocomposites

Figure (6) shows FT-IR spectrum of (nCePf) / polyaniline nanocomposites (I, II, III and IV), respectively, (putting in consideration slight shifts of some bands as well as the low band intensity of composite (I) in the region $1723\text{--}1430\text{ cm}^{-1}$ this was attributed to low loading of PANI in the composite). Broad band centered around 3442 cm^{-1} , is due to OH groups symmetric stretching of H_2O superimposed with the N-H stretching of aromatic amines (expected at the range $3500\text{--}3280\text{ cm}^{-1}$). Very small band around 1620 cm^{-1} is related to H-O-H bending, and sharp broad band, centered at

(~ 1048 or $\sim 1005\text{ cm}^{-1}$) corresponds to phosphate groups vibration. Small band at 2924 cm^{-1} corresponds to C-H bonds. The presence of two bands in the range of 1490 , 1303 cm^{-1} is assigned to the non-symmetric C_6 ring stretching modes. However the higher frequency vibration at 1573 cm^{-1} has a major contribution from the quinoid rings ($\text{C}=\text{C}$ stretching vibration of quinoid ring), while the lower frequency mode at 1490 cm^{-1} depicts the presence of benzenoid rings ($\text{C}=\text{C}$ stretching vibration of benzenoid ring). Thus FTIR spectra confirms the formation of composites [54, 55].

9. TGA of Nanocomposites

Thermogram (TG/DTA) of (nCePf) / PANI, composite (I), is shown in Figure (7). The weight loss up to 200°C (9.61%) is due to the removal of external water molecules. The weight loss occurs between $200\text{--}700^\circ\text{C}$ corresponds to the decomposition of polyaniline and condensation of P-OH groups of the inorganic material to pyrophosphate (CeP_2O_7). The POH groups condensation found to superimpose with that of polyaniline decomposition. Thermal decomposition was accompanied with four endothermic peaks at 70 , 190 , 300 , 350°C . The total weight loss found to be equal to 21.899%. From thermal and elemental (C, H, N) analysis the resultant product was formulated as: $\text{Ce}(\text{HPO}_4)_2 (\text{PANI})_x \cdot 2.2\text{ H}_2\text{O}$, where $x = 7.82\%$.

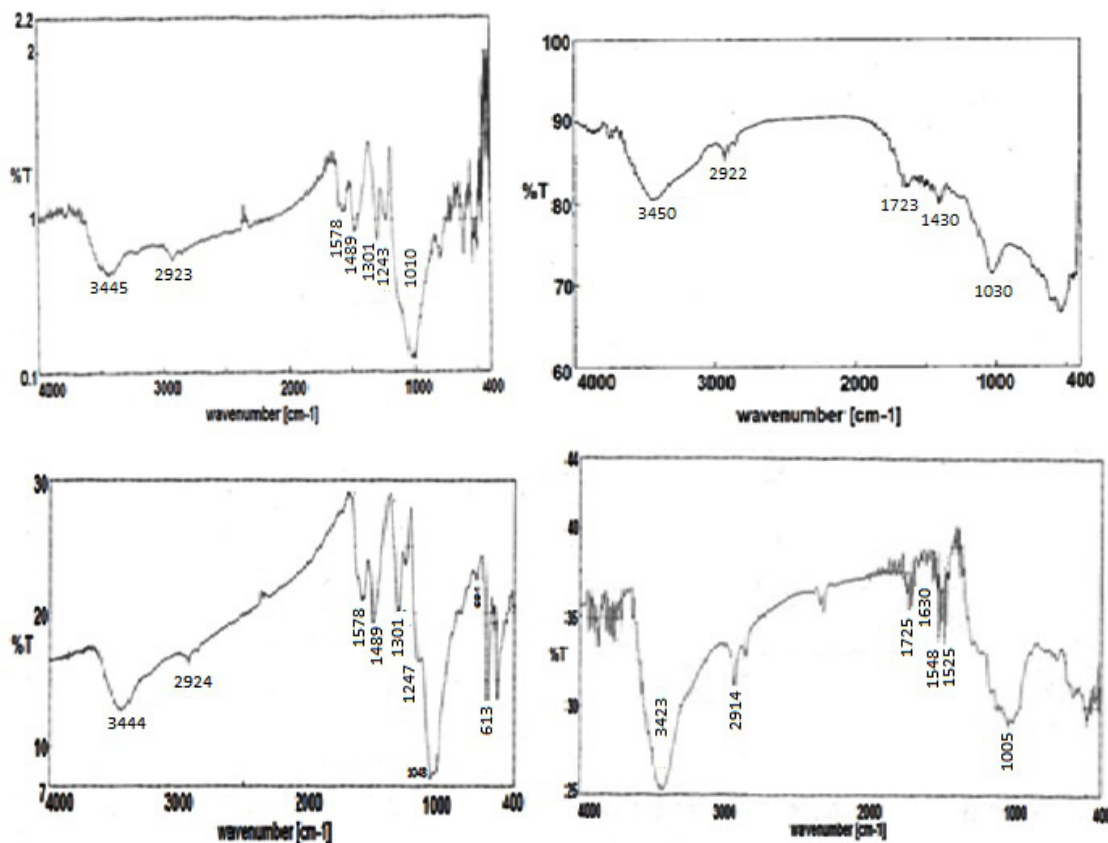


Figure 6. FT-IR of (nCePf)/PANI nanocomposites(I-IV) respectively

Thermogram (TG/DTA) of (nCePf) / PANI, composite (IV), is also shown in Figure (7), in respective manner. The weight loss up to 200°C (8.93%) is due to the removal of external water molecules. The weight loss occurs between 200-800°C corresponds to the decomposition of polyaniline and condensation of P-OH groups of the inorganic material to pyrophosphate (CeP_2O_7). The POH groups condensation found to superimpose with that of polyaniline decomposition. Thermal decomposition was accompanied with three endothermic peaks at 90, 240, 560°C. The total weight loss found to be equal to 32.203%. From thermal and elemental (C, H, N, S) analysis the resultant product was formulated as: $\text{Ce}(\text{HPO}_4)_2 (\text{PANI})_x (\text{SO}_4^{2-})_y \cdot 2.51 \text{H}_2\text{O}$, where $x = 17.71\%$ in weight, $y = 2.05\%$ in weight.

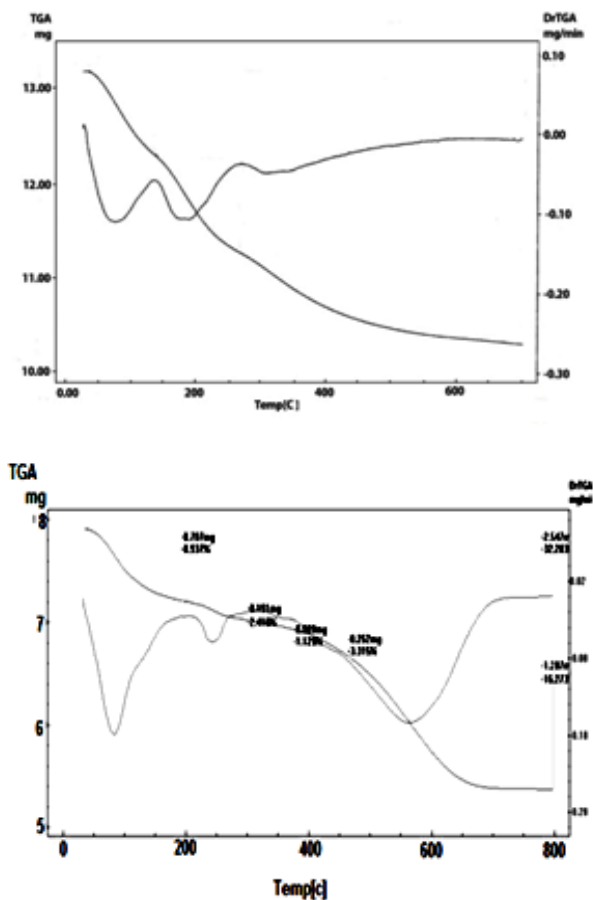


Figure 7. TG/DTA of (nCePf)/polyaniline composites(I & IV) respectively

3. Conclusions

Novel supported nanofibrous cerium phosphate/polyaniline nanocomposite membranes were prepared via in-situ chemical oxidation of the monomer that was promoted by the reduction of Ce(IV) ions present in the inorganic matrix. The presence of Ce(IV) ions allows redox reactions necessary to oxidative polymerization to occur. The self-supported sheet integrity is preserved. This result is interesting because the shape integrity makes the building of molded conducting device possible. The formulation of the

fibrous cerium phosphate nanocomposites was supported by thermal, elemental (C, H, N) analysis, FT-IR spectrum and SEM. The type of polyaniline composite and the degree of polymer loading found to be dependant on variation of method of the preparation and time aging. Polyaniline (PANI) is generally considered one of the most intriguing material for sensing devices due to its environmental stability, high conductivity. Thus by utilizing the beneficial properties of both cerium phosphate and PANI, their hybrid can be considered as novel conducting inorganic-organic composites, ion exchangers, solid acid catalyst and as sensors [10, 55, 56].

ACKNOWLEDGEMENTS

To Department of Chemistry, Faculty of Science, Tripoli University, for providing facilities for this research, to Geol. Adel Bayuomi, Institute of Mineral Resources, Cairo, Egypt for providing facilities for elemental CHNS, FT-IR, TGA, SEM analysis.

REFERENCES

- [1] Street, G. B.; Clarke, T. C., Conducting Polymers: A Review of Recent Work IBM J. Res. Dev. 25(1981)51-57.
- [2] T. J. Skotheim, J. R. Reynolds, Eds. Handbook of Conducting Polymers, Third Ed., CRC Press, Boca Raton, USA (1997).
- [3] H. S. Nalwa, Ed. Handbook of Organic Conductive Molecules and Polymers, Vol. 2, John Wiley, Chichester, UK (1997).
- [4] MacDiarmid, A. G., A novel role for organic polymers, Angewandte Chemie International Edition, 40 (2001)2581.
- [5] Heeger, A., Semiconducting and metallic polymers: The fourth generation of polymeric materials, Reviews of Modern Phys, 73(2001) 681.
- [6] Freund, M.S. and Deore, B.A. Edts., Self Doped Conducting Polymers, John Wiley and Sons (2007).
- [7] Lange, U., Roznyatovskaya, N. V., Mirsky, Vladimir M., Conducting polymers in chemical sensors and arrays, Analytica Chimica Acta vol. 614 issue (2008).
- [8] Kucheldorf, H. R., Luken, O. and Swift, G., Hand book of polymer synthesis, 2nd Edit., CRC Press (2010).
- [9] Banerjee, S. and Tayagi, A.K. Edts, Functional Materials: Preparation, Processing and Applications, Elsevier (2012).
- [10] Lange, U., Roznyatovskaya, N. V., Mirsky, Vladimir M., Conducting polymers in chemical sensors and arrays, Analytica Chimica Acta vol. 614 issue (2008).
- [11] Sapurina, I. Y. and Shishov, M. A., New polymers for special applications, Ailton De Souza Gomes, Ed. (2012).
- [12] Cai, Z., Grang, M. and Tang, Z., Novel battery using conducting polymers: and polyaniline, J. of Mater. Sci., 39 (2004) 4001.

- [13] Ćirić-Marjanović, G., Polyaniline nanostructures, in nanostructured conductive polymers, (Eftekhari, A. ed., John Wiley & Sons, Ltd (2010).
- [14] Vernitskaya, T.V. Tat'yana, V.V. and Efimov, Polypyrrole: a conducting polymer; its synthesis, properties and applications, *Russ. Chem. Rev.*, (1997)66 443.
- [15] Feng-Hao Hsu, F.H and, Wu, T. M., In situ synthesis and characterization of conductive polypyrrole/graphene composites with improved solubility and conductivity, *Synthetic Metals*, 162(2012) 682–687.
- [16] Fajan, A. and Been, B., structural and optical properties of polyindole manganese oxide nanocomposite, *Indian J. of Adv. in Chem. Sci.*, 2 (2013) 95-96.
- [17] Shukla, S.K., Bharadvaja, A., Tiwari, A., Parashar, G.K., and Dubey, G. C., Synthesis and characterization of highly crystalline polyaniline film promising for humid sensor., *Adv. Mat. Lett.*, 1(2010) 129-134.
- [18] Chiang, J.C., MacDiarmid, A. G., Protonic acid doping of the emeraldine form to the metallic regime, *Synthetic Metals*, 1(1986) 193.
- [19] Ito, S., K. Murata, K. , S. Teshima, S., R. Aizawa, R., Y. Asako, Y, Takahash K. and Hofmann, B. M., Simple synthesis of water-soluble conducting polyaniline, *Synthetic Metals*, 96(1998) 161.
- [20] Colak, N. and Sokman, B., Doping of Chemically synthesized polyaniline, *Des Monomer Polym.*, 3 (2000) 181-189.
- [21] MacDiarmid, Alan G, A novel role for organic polymers, *Angewandte Chemie International Edition*, 40 (2001)2581.
- [22] Heeger, A., Semiconducting and metallic polymers: The fourth generation of polymeric materials, *Reviews of Modern Phys*, 73(2001)681.
- [23] Virji, S., H., Jiaying; Kaner, R. B.; Weiller, B. H., Polyaniline nanofiber gas Sensors: examination of response mechanisms, *Nano Letters*, 4(2004) 491.
- [24] Ansaci, R.M. and Kervani, F.V., Polyaniline conducting electroactive polymers: thermal and environmental stability studies, *J.Chem.*, 3(2006) 202-217.
- [25] Li, X., Oxidative polymerization of aniline using NaClO₂ as an oxidant, *Mater. Lett.*, 61,(2007) 2011.
- [26] Wessling, B., New insight into organic metal polyaniline morphology and structure, *Polym.*, 2(2010) 786.
- [27] Sapurina, I. Y. and Shishov, M. A., New polymers for special applications, Ailton De Souza Gomes, Ed. (2012).
- [28] Chen.y, Li, Y., Yip, M., Taid, N., Electromagnetic interference shielding efficiency of polyaniline composites filled with graphene decorated with metallic nanoparticles, *J. Composite Sci. Tech*, 80(2013)80.
- [29] Sharma, D., Kaith, B.S. and Rajput, J, Single step in situ synthesis and optical properties of polyaniline/ZnO nanocomposites, *The Sci. World J*, Volume 2014 (2014), Article ID 904513, 13 pages.
- [30] Pandiselvi, K., Manikumar, A. and Thambidurai, S., Synthesis of novel polyaniline/MgO composite for enhanced adsorption of reactive dye, *Polymer Composites*, 35(2014)351.
- [31] Clearfield, A., *Inorganic Ion Exchange Materials*, Bocca Raton, CRC Press Fl.. (1988).
- [32] Colon, J.L., Diaz, A. and Clearfeild, A., Nanoencapsulation of insulin into zirconium phosphate for oral delivery applications, *Biomacro Molecules*, 9 (2010) 2465.
- [33] Diaz, A., Saxena, V., Gunzalez, J., David, A., Casanas, B., Carpenter, C., Batteas, J.D., Colon, J., Clearfield, A. and Hussain, M.D., Zirconium phosphate nano-platelets: a novel platform for drug delivery in cancer therapy, *Chem. Commun.*, 48 (2012) 1754.
- [34] Tushato, M., Danjo, M., Baba, Y., Murakom, M. and Nana, H., Preparation and chemical properties of a novel layered cerium(iv) phosphate, *Bulletin of Chem.Soc Jap.*, 70 (1997) 143.
- [35] Salvado, M.A., Pertierra, P., Tropajo, C. and Garcia, G.R., Crystal structure of cerium(iv) bis(hydrogen phosphate derivative, *J. Am. Chem. Soc.*, 129(2007)10970.
- [36] Alberti, G. and Costantino, U., Recent progress in the field of synthetic inorganic exchanger having a layered and fibrous structure, *J. Chromatogr.*, 102 (1974)5.
- [37] Parangi, T., Wani, P. and Chudasama, U., Synthesis, characterization and application of cerium phosphate, *Desalination and Water Treatment*, 38 (2012)126.
- [38] Romano, R. and Alves, O.S., Fibrous cerium(iv) phosphate host of weak and strong Lewis bases, *Inclus. Phenom. and Macrocyclic*, 51 (2005) 211.
- [39] Casciola, M., Costantino, U. and D'amico, S., ac Conductivity of cerium (iv) phosphate in hydrogen form, *Solid State Ionics*, 28 (1988) 617.
- [40] Varsheny, K.G., Tayal, N., Gupta, P, Agrawal, A and Drabik, M., Synthesis, ion-exchange and physical studies on polystyrene cerium(iv) phosphate hybrid fibrous ion exchanger , *Ind. J. of Chem .* 43 (2004) 2586.
- [41] Metwally, S.S., El-Gammal, B., Ali, H.F. and Abo-EL-Enein, S.A., Removal and separation of some radionuclides by poly-acrylamide based Ce(iv) phosphate, *Separation Sci. and Tech.*, 46 (2011)11.
- [42] Alberti, G., Casciola, M., Captani, D., Donnadio, A., Narducci, R., Pica, M. and Sganappa, M., N., Novel nafion-zirconium phosphate composite membranes with enhanced stability of proton conductivity, *Electrochimica Acta*, 52 (2007)8125.
- [43] Yang, Y., Liu, C. and Wen, H., Preparation and properties of polyvinyl alcohol / exfoliated α -zirconium phosphate, *Polym. Test.*, 28 (2009)185.
- [44] Nagarale, R.K., Shin, W and Singh, P.K., Progress in ionic organic-inorganic composite membranes for fuel cell application, *Polym. Chem.*, 1 (2010) 388.
- [45] S. K. Shakshooki, S.K, Elejmi, A.A., Khalfulla, A. M. And. Elfituri. S. S., *Int. Conf. on Mater. Imperative*, pp49-70 (CD). Cairo, Egypt, 29/11-2/12 (2010) .
- [46] Feng, Y., He, W., Zhang, X., Jia, X. and Zhao, H., The preparation of nanoparticle zirconium phosphate, *Mater.. Letters*, 61 (2007) 3258.

- [47] Bao, C., Gao, Y., Song, L., Lu, H., Yuan, B., and Hu, Y., Facile synthesis of poly(vinyl alcohol)/ α -titanium phosphate nanocomposite with markedly enhanced properties, *Ind. Eng. Chem. Res.*, 50 (2011)11109.
- [48] Verissimo, C. and Alves, O.L, Preparation of the conducting nanocomposites using molded inorganic matrix: fibrous cerium phosphate as a self supported pyrrole polymerization, *J. Mater. Sci.*, 13 (2003)1378-1383.
- [49] Salomaki, M., Rasanen, M., Leiro, J., Huti, T., Tenho, M., Lukkari, J. and Kankari, J., Oxidative inorganic multilayer for polypyrrole film generation, *Adv. Funct. Mater.*, 20 (2010) 2140-2147.
- [50] Rosenthal, G.L., Caruso, J., and Stone, S.G., Polymerization of aniline on copper zirconium phosphate and mixed phosphonate, *Polyhedron*, 13(1999)1311-1314.
- [51] Goncalves, A.B., Mangrich, A.S. and Zarbin, A.J.G., Polymerization of pyrrole between the layers of tin(IV) bisphosphate, *Synthetic metals*, 114(2000)119-124.
- [52] Shakshooki, S.K., El-Akari, F.A., El-Fituri, M.S. and El-Fituri, S.S., Fibrous cerium(IV) hydrogen phosphate membrane self supported benzimidazole polymerization agent, *Adv. Mater. Res.*, 856 (2014) 3.
- [53] Shakshooki, S.K., Nano Fibrous cerium (IV) hydrogen phosphate membrane self supported indole polymerization agent, *J. Chem. Chem. Eng.*, 8(2014) 378-384.
- [54] Parveen, A. and Roy, A., Effect of morphology on thermal stability of core-shell polyaniline/TiO₂, *Adv.Mat.Lett.*, 4(2013) 696-701.
- [55] Khuspe, G.D., Navale, S.T., Bandgar, Sakhare, D. K., Chugle, M.A., and Patil, V.B., SnO₂ nanoparticles-modified Polyaniline Films as Highly Selective, Sensitive, Reproducible and Stable Ammonia Sensors, *Electron. Mater. Lett.*, 10, (2014) 191-197.
- [56] R. Jain, D. C. Tiwari and S.Shrivastan, Sensitive voltmmetric sensor based on synergistic effect of polyaniline and zirconia nanocomposite film, *J. of Elect. Chem. Soci.*, 161(2014) B38-B44.