

# 1-Butyl-3-Methylimidazolium Tetrafluoroborate ([BMIM][BF<sub>4</sub>]) and Hexafluorophosphate ([BMIM][PF<sub>6</sub>]) Mediated Synthesis of Bicyclic Lactones

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**Abstract** The transition toward sustainable solvent systems is essential for alleviating the environmental and safety risks associated with traditional organic synthesis methodologies. This study investigates the sodium borohydride (NaBH<sub>4</sub>) reduction of (1S,2R,3S)-3-methyl-2-(nitromethyl)-5-oxocyclopentylacetic acid using two ionic liquids—1-butyl-3-methylimidazolium tetrafluoroborate ([BMIM][BF<sub>4</sub>]) and hexafluorophosphate ([BMIM][PF<sub>6</sub>])—as environmentally benign alternatives to ethanol. The reduction conducted at room temperature in [BMIM][PF<sub>6</sub>] yielded a *syn:anti* product ratio of 74:26, whereas [BMIM][BF<sub>4</sub>] resulted in a ratio of 69:29, in stark contrast to ethanol's significantly lower *syn* selectivity of 21:69, accompanied by the formation of an undesired byproduct identified as an amine from the reduction of the nitro group. At a temperature of 0°C, [BMIM][PF<sub>6</sub>] further improved selectivity to 83:17, demonstrating enhanced control over stereochemistry. These outcomes are attributed to the distinctive solvation environment offered by ionic liquids, which inhibit NaBH<sub>4</sub> decomposition, enhance substrate solubility, and stabilize reactive intermediates. Furthermore, these solvents facilitate the simplification of product isolation, reduce waste generation, and eliminate risks associated with solvent evaporation. The findings highlight the potential of ionic liquids to replace volatile organic solvents, thereby fostering safer, more efficient, and environmentally sustainable practices in organic synthesis.

**Keywords** Green Chemistry, Ionic Liquids, NaBH<sub>4</sub> Reduction, Sustainable Synthesis, [BMIM][BF<sub>4</sub>], [BMIM][PF<sub>6</sub>]

## 1. Introduction

The utilization of sodium borohydride (NaBH<sub>4</sub>) as a chemoselective reducing agent is well-documented, particularly in the conversion of carbonyl functionalities to their corresponding alcohols [1]. These reductions are typically conducted in protic solvents such as ethanol or methanol, which facilitate the solubilization of both the substrate and the reducing agent [2]. However, these solvents present environmental and safety challenges, including flammability, volatility, and the generation of hazardous waste streams [3].

In response to these issues, ionic liquids (ILs) have emerged as environmentally benign alternatives. Comprising bulky organic cations and weakly coordinating anions, ILs exhibit properties such as negligible vapor pressure, high thermal stability, and non-flammability [4,5]. Their tunable solvation environments and structural diversity make them suitable mediums for a variety of chemical transformations,

rendering them attractive candidates for incorporation within green synthetic methodologies [6,7]. Numerous studies have indicated that ILs not only create a safer reaction environment but also enhance selectivity, catalytic efficiency, and sustainability across a spectrum of reactions, including reductions utilizing hydride donors [8,9]. In these reactions, ILs improve substrate solubility, stabilize reactive intermediates, and mitigate decomposition pathways that are typically observed with NaBH<sub>4</sub> in protic solvents [10,11]. Moreover, certain ionic liquids have been shown to influence stereoselectivity, a phenomenon attributed to their polarizable and highly organized solvation shells, which can preferentially favor the formation of one diastereomer over another [12].

This study investigates the reduction of (1S,2R,3S)-3-methyl-2-(nitromethyl)-5-oxocyclopentylacetic acid through the application of NaBH<sub>4</sub> in two hydrophobic ionic liquids, [BMIM][BF<sub>4</sub>] and [BMIM][PF<sub>6</sub>], compared to ethanol. The objective is to quantify variations in stereoselectivity, yield, and efficiency while also underscoring the advantages of ionic liquids as sustainable solvents for fine chemical and pharmaceutical synthesis. The broader implication of this work is the demonstration of ionic liquids as feasible

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substitutes for volatile organic solvents in industrial-scale reductions, thus contributing to greener and safer chemical manufacturing practices.

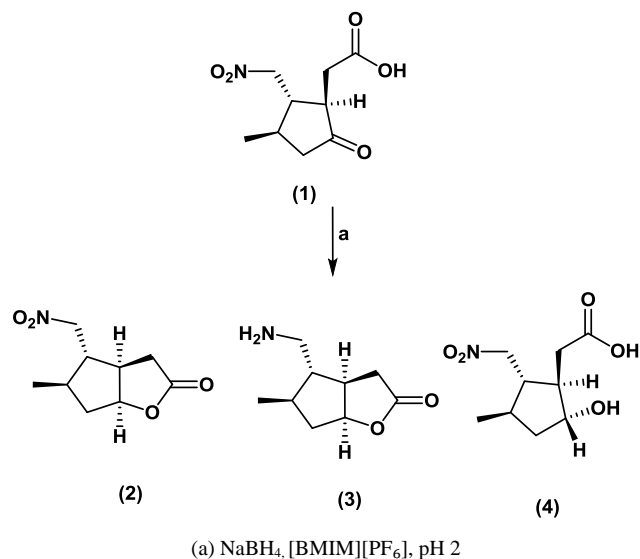
## 2. Results and Discussion

The reduction of (1S,2R,3S)-3-methyl-2-(nitromethyl)-5-oxocyclopentylacetic acid (**1**) at 0 °C utilizing NaBH<sub>4</sub> in ethanol, [BMIM][BF<sub>4</sub>], and [BMIM][PF<sub>6</sub>] yields a diastereomeric mixture of *syn* and *anti*-alcohols. However, the selectivity for the *syn* isomer, which readily undergoes cyclization, exhibited considerable variability. [BMIM][PF<sub>6</sub>] produced a *syn:anti* ratio of 83:17, whereas [BMIM][BF<sub>4</sub>] resulted in a 73:27 ratio. In contrast, the reduction in ethanol favored the *anti*-isomer, yielding a *syn:anti* ratio of 25:65. These results underscore the significance of ionic solvents in facilitating the formation of the desired bicyclic compound.

When the reaction was repeated at room temperature for 1 hour in ethanol, TLC analysis showed the appearance of a new product. The solvent was removed under reduced pressure and the residue was subjected to gradient column chromatography, affording compounds (**2**), (**4**), and (**3**) in a ratio of 21:69:10, respectively. Their masses, IR, and <sup>1</sup>H NMR spectra fully supported the proposed structures (Scheme 1). These findings emphasize the advantageous role of ionic liquids in enhancing reaction efficiency compared to traditional protic solvents such as ethanol. It is posited that the improved yields observed in [BMIM][BF<sub>4</sub>] and [BMIM][PF<sub>6</sub>] are attributable to several intrinsic properties of ionic liquids. Ethanol, as a protic solvent, can partially decompose NaBH<sub>4</sub> through protonation, leading to decreased availability of the reducing agent [13,14]. In contrast, ionic liquids offer a non-protic, highly stable environment that prevents premature decomposition of NaBH<sub>4</sub> and promotes more efficient hydride transfer [15,16].

Additionally, ionic liquids, particularly [BMIM][PF<sub>6</sub>], enhance substrate solubility and enable improved interaction between NaBH<sub>4</sub> and the ketone due to their unique solvating ability and polarizable microenvironments [17,18]. Ethanol, by comparison, often suffers from solubility limitations and competitive hydrogen bonding, which can diminish reduction efficiency. Moreover, the higher viscosity of ionic liquids slows diffusion, enabling a controlled release of hydride ions and minimizing side reactions [19,20]. Literature reports further suggest that ionic liquids stabilize transition states, thereby increasing reaction selectivity and suppressing undesired byproducts [21,22]. In 2014, Ratti demonstrated that ionic liquids stabilize reactive intermediates and enhance NaBH<sub>4</sub> reduction yields relative to conventional solvents [23]. Zhao *et al.* (2005) showed that ILs with hydrophobic anions, such as PF<sub>6</sub><sup>-</sup>, provide a superior reaction environment compared to their hydrophilic counterparts, explaining the performance difference between [BMIM][BF<sub>4</sub>] and [BMIM][PF<sub>6</sub>] [24]. Welton has also noted that ionic liquids can increase electron density at reactive centers, further promoting efficient reductions [25].

We envisaged the observed variability in stereoselectivity between ethanol and ionic liquids likely stems from differences in solvation dynamics and microenvironmental effects. Ethanol, a small protic solvent, exhibits limited stereochemical control, whereas ionic liquids offer a more structured and directional environment that can significantly influence the hydride delivery mechanism from sodium borohydride (NaBH<sub>4</sub>). The interactions between the cation and substrate, along with the nature of the anion (e.g., PF<sub>6</sub><sup>-</sup> versus BF<sub>4</sub><sup>-</sup>), likely stabilize specific transition states, thereby promoting *syn*-selectivity. This phenomenon becomes increasingly pronounced at lower temperatures due to diminished thermal motion and enhanced ordering within the ionic liquid medium.



**Scheme 1.** Stereoselectivity in NaBH<sub>4</sub>-Mediated Ketone Reduction

**Table 1.** Stereoselectivity in NaBH<sub>4</sub>-Mediated Ketone Reduction

Solvent	Temp. (°C)	<i>syn</i> -( <b>2</b> ): <i>anti</i> ( <b>4</b> ) : ( <b>3</b> )
Ethanol	RT	21:69:10
	0 °C	25:65
[BMIM]PF <sub>6</sub>	RT	74:26
	0 °C	83:17
[BMIM]BF <sub>4</sub>	RT	69:29
	0 °C	73:27

The IR spectrum revealed a stretching at 1710 cm<sup>-1</sup> and this is not expected for a 5-membered lactone, but it is well known that hydrogen bonding at the oxygen of a carbonyl group reduces the frequency of C=O stretching by lowering the bond order, in some cases below 1700cm<sup>-1</sup>. We therefore rationalized this to have arisen presumably because of intermolecular hydrogen bonding between the primary amine and the lactone. There was also the appearance of an absorption due to amine group at 3402 cm<sup>-1</sup>.

## 3. Conclusions

The findings of this study strongly support the use of ionic liquids as additional good solvent option for NaBH<sub>4</sub> reductions

in advanced organic chemistry experiments. Replacing ethanol with [BMIM]BF<sub>4</sub> and [BMIM]PF<sub>6</sub> resulted in higher reaction yields, improved stereoselectivity, and enhanced reaction control, demonstrating the transformative potential of ionic liquids in organic synthesis. Notably, the reaction in [BMIM]PF<sub>6</sub> exhibited the highest efficiency, yielding an 83:17 *syn:anti* ratio, while [BMIM]BF<sub>4</sub> produced a 73:27 *syn:anti* ratio, both significantly outperforming ethanol. The advantages of ionic liquids over ethanol in NaBH<sub>4</sub> reductions can be attributed to several key factors: minimized NaBH<sub>4</sub> decomposition—unlike ethanol, which can lead to premature hydrolysis of NaBH<sub>4</sub>, ionic liquids provide a more stable reaction environment, ensuring greater hydride availability throughout the reaction; enhanced solubility and reaction kinetics—the unique solvation properties of ionic liquids facilitate better dissolution of both reactants and intermediates, leading to improved reduction efficiency and selectivity; increased stereoselectivity favoring the *syn* product—the ionic medium influences transition-state stabilization, resulting in a higher proportion of the desired *syn* isomer, particularly evident in [BMIM]PF<sub>6</sub> (83:17 *syn:anti* ratio), compared to ethanol, where the ratio was significantly lower; and greater sustainability and safety—unlike ethanol, which is flammable, volatile, and generates hazardous waste, ionic liquids eliminate solvent evaporation, reduce environmental impact, and offer safer handling conditions in laboratory settings. Additionally, integrating ionic liquids into undergraduate laboratory experiments aligns with the Principles of Green Chemistry, offering a sustainable, high-efficiency alternative to traditional organic solvents. By adopting ionic liquids, students gain hands-on experience with modern, eco-friendly solvent systems while reinforcing concepts of stereoselectivity, solubility effects, and solvent-dependent reaction dynamics. Moreover, the ability to recycle and reuse ionic liquids further underscores their potential as viable replacements for hazardous volatile solvents in teaching and research laboratories.

### 3.1. Experimental Techniques

Commercial reagents were obtained from Thermo Scientific or VWR and were used directly as supplied or purified prior to use, following established guidelines.

Non-aqueous reagents were transferred under nitrogen using a syringe. Organic solutions were concentrated under reduced pressure on a Büchi rotary evaporator using a water bath or were evaporated under a sand bath. Thin-layer chromatography (TLC) was performed on Merck aluminum-backed plates coated with 0.2 mm silica gel 60-F. Visualization of the developed chromatogram was conducted using UV fluorescence quenching at 254 nm.

<sup>1</sup>H and <sup>13</sup>C NMR spectra were recorded on a Jeol JNM-ECZ (400 MHz for protons). Data for <sup>1</sup>H NMR are reported as follows: chemical shift (δ-ppm), multiplicity (s = singlet, d = doublet, t = triplet, q = quartet, m = multiplet), integration, and coupling constant (Hz). Data for <sup>13</sup>C NMR spectra are reported in terms of chemical shift (ppm) downfield from TMS.

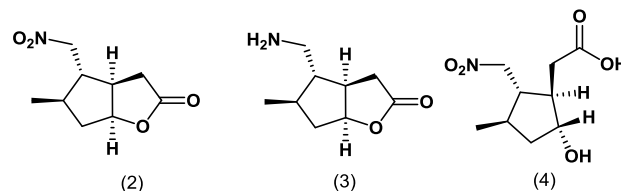
IR spectra were recorded on a Bruker Alpha spectrometer using direct ATR application. All absorptions are reported in terms of frequency of absorption (cm<sup>-1</sup>).

### 3.2. Experimental

**(3a*R*,4*S*,5*R*,6a*S*)-5-methyl-4-(nitromethyl)-hexahydrocyclopenta[*b*]furan-2-one (2)**

**(3a*R*,4*S*,5*R*,6a*S*)-4-(aminomethyl)-5-methyl-hexahydrocyclopenta[*b*]furan-2-one (3)**

**2-(1*R*,2*S*,3*R*,5*R*)-5-hydroxy-3-methyl-2-(nitromethyl)cyclopentyl)acetic acid (4)**



### Method A

To a stirred solution of [(1*S*,2*R*,3*S*)-3-methyl-2-(nitromethyl)-5-oxocyclopentyl]acetic acid (**1**) (2.00 g, 9.95 mmol, 1.00 equiv) in [BMIM]PF<sub>6</sub> (10 mL) within an Erlenmeyer flask, sodium borohydride (NaBH<sub>4</sub>) was added at room temperature (489 mg, 12.94 mmol, 1.30 equiv) in small portions over a period of 15 minutes. The resulting reaction mixture was stirred for an additional 30 minutes and subsequently poured into ice water (10 mL). Dilute hydrochloric acid (5%) was then added until a pH of 2 was achieved, at which point thin-layer chromatography (TLC) analysis indicated the formation of a new product. The aqueous layer was extracted with ether (4 x 10 mL), the extract was dried over magnesium sulfate (MgSO<sub>4</sub>), and concentration under reduced pressure afforded a mixture (1.82 g, 91% yield). Column chromatography on silica, eluting with a hexane:ether mixture (5:1), yielded a mixture of compounds: (**2**), (**3**), and (**4**) (1.35 g, 74%; 0.22 g, 12%; 0.25 g, 14%, respectively).

(**2**) was obtained as a colorless oil (3.9 g, 85% yield);  $\nu_{\text{max}}$  (thin film) 2968, 2931, 1769, 1550, 1381, 1172;  $\delta_{\text{H}}$  (400 MHz, CDCl<sub>3</sub>) 4.92 (1H, q, *J* 5.2 Hz, *J* 2.2 Hz, CHO), 4.60 (1H, dd, *J* 7.5 Hz, *J* 4.7 Hz, CH<sub>2</sub>NO<sub>2</sub>), 4.37 (1H, dd, *J* 8.4 Hz, *J* 3.8 Hz, CH<sub>2</sub>NO<sub>2</sub>), 2.78 (1H, dd, *J* 9.8 Hz, *J* 8.4 Hz, *J* 1.6 Hz, CH<sub>2</sub>C=O), 2.73 (1H, ddd, *J* 5.6 Hz, *J* 1.9 Hz, CHCHO), 2.54-2.47 (1H, m, CH<sub>2</sub>CHO), 2.44 (1H, dd, *J* 15.9 Hz, *J* 1.9 Hz, CH<sub>2</sub>C=O), 2.06-2.03 (1H, m, CHCH<sub>2</sub>NO<sub>2</sub>), 1.82-1.78 (1H, m, CHCH<sub>3</sub>), 1.62-1.54 (1H, m, CH<sub>2</sub>CHO), 1.11 (3H, d, *J* 6.5 Hz, CH<sub>3</sub>);  $\delta_{\text{C}}$  (100 MHz, CDCl<sub>3</sub>) 176.2, 83.1, 77.4, 50.5, 43.5, 40.6, 37.4, 34.0, 17.4; *m/z* (C.I) 200 (MH<sup>+</sup>, 100%), 182 (54%), 153 (58%), 131 (60%), 108 (40%), 82 (21%) C<sub>9</sub>H<sub>14</sub>N<sub>4</sub>, requires 200.0924, found 200.0916.

(**3**) was obtained as a colorless oil (196 mg, 12% yield) with the following spectral data:  $\nu_{\text{max}}$  (thin film/cm<sup>-1</sup>) 3402, 2957, 1718, 1549;  $\delta_{\text{H}}$  (400 MHz, CDCl<sub>3</sub>) 4.49 (2H, d, *J* 6.6 Hz, CH<sub>2</sub>NH<sub>2</sub>), 4.04 (1H, ddd, *J* 5.0 Hz, *J* 2.9 Hz, *J* 2.8 Hz, CHOH), 2.52 (1H, dd, *J* 11.8 Hz, *J* 4.9, CH<sub>2</sub>C=O), 2.38 (1H, dd, *J* 9.7 Hz, *J* 7.3 Hz, *J* 2.3 Hz, CH<sub>2</sub>C=O), 2.11-2.03 (2H, m, CHCH<sub>2</sub>C=O, CHCH<sub>3</sub>), 1.93-1.83 (2H, m, CHCH<sub>2</sub>NH<sub>2</sub>,

CH<sub>2</sub>CHO), 1.59–1.54 (1H, m, CH<sub>2</sub>CHO), 1.06 (3H, d, *J* 6.0 Hz, CH<sub>3</sub>CH); δC (100 MHz, CDCl<sub>3</sub>) 177.0, 78.7, 77.4, 49.7, 48.8, 41.1, 37.5, 35.6, 19.3; m/z (C.I) 153 (M<sup>+</sup>-NH<sub>2</sub>, 96%), 139 (20%), 132 (28%), 109 (21%), 107 (51%) C<sub>9</sub>H<sub>15</sub>N<sub>2</sub>, requires 153.0916, found 153.0912.

(4) was obtained as a white powder (81 mg, 4% yield) with a melting point of 65–67°C; ν<sub>max</sub> (KBr/cm<sup>-1</sup>) 3461, 2967, 2936, 1733, 1553, 1378; δH (400 MHz, CDCl<sub>3</sub>) 4.56 (1H, dd, *J* 8.8 Hz, *J* 3.6 Hz, CH<sub>2</sub>NO<sub>2</sub>), 4.29 (1H, dd, *J* 8.3 Hz, *J* 3.6 Hz, CH<sub>2</sub>NO<sub>2</sub>), 3.41 (1H, m, CHOH), 2.42–2.36 (2H, m, CH<sub>2</sub>C=O), 2.01 (1H, m, CHCHOH), 1.86–1.83 (1H, m, CH<sub>2</sub>CHOH), 1.78–1.74 (1H, m, CHCH<sub>2</sub>NO<sub>2</sub>), 1.68–1.64 (1H, m, CHCH<sub>3</sub>), 1.58–1.56 (1H, m, CH<sub>2</sub>CHOH), 1.08 (3H, d, *J* 6.4 Hz, CH<sub>3</sub>); δC (100 MHz, CDCl<sub>3</sub>).

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